

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

*The changes in energy for these reactions are the same, but opposite in sign.*

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Pogil Bond Energy Key

**Mark Steyvers**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## Pogil Bond Energy Key:

Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more **Comprehensive Handbook of Chemical Bond Energies** Yu-Ran Luo, 2007-03-09

Understanding the energy it takes to build or break chemical bonds is essential for scientists and engineers in a wide range of innovative fields including catalysis nanomaterials bioengineering environmental chemistry and space science Reflecting the frequent additions and updates of bond dissociation energy BDE data throughout the literature the Comprehensive Handbook of Chemical Bond Energies compiles the most recent experimental BDE data for more than 19 600 bonds of 102 elements The author organizes the data by bond type functional group bond order bond degree molecular size and structure for ease of use Data can also be located using the Periodic table The book presents data for organic molecules biochemicals and radicals as well as clusters ions hydrogen and surface bonded species van der Waals complexes isotopic species and halogen clusters complexes It also introduces entirely new data for inorganics and organometallics The final chapter summarizes the heats of formation for atoms inorganic organic radicals and monoatomic ions in the gas phase The Comprehensive Handbook of Chemical Bond Energies offers quick access to experimental BDE data in the most inclusive well organized and up to date collection available today CHEMICAL BONDS AND BOND ENERGY. RT.

SANDERSON, 1971 **Bond Orders and Energy Components**, 2013 Bond Energy Study of Group 6 Metal-olefin Complexes  $M-(CO)_5(C_2H_4-nX_n)$ , X Sheryl L. Johnson, Illinois State University. Department of Chemistry, 2013 *Bond Orders and Energy Components* István Mayer, 2016-10-14 While modern computational methods can provide us with the wave function of a molecule in numerical form most computer programs lack the sophisticated tools needed to extract chemical concepts from these wave functions Saving researchers vast time and potential confusion this volume collects and organizes those validated tools currently scattered throughout the literature and details their application It provides immediate access for those needing to calculate such critical factors as bond order and valence indices and atomic and diatomic contributions to molecular energy Supporting material is available for download from the authors continually updated website **Energy Refinements to Glass Structure** Patrick Joseph Hood, 1985 Chemistry Richard S. Moog, John J. Farrell, 2014-01-13 Chemistry A Guided Approach 6th Edition follows the underlying principles developed by years of research on how readers learn and draws on testing by those using the POGIL methodology This text follows inquiry based learning and correspondingly emphasizes the underlying concepts and the reasoning behind the concepts This text

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## **Table of Contents Pogil Bond Energy Key**

1. Understanding the eBook Pogil Bond Energy Key
  - The Rise of Digital Reading Pogil Bond Energy Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Bond Energy Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pogil Bond Energy Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Bond Energy Key
  - Personalized Recommendations
  - Pogil Bond Energy Key User Reviews and Ratings
  - Pogil Bond Energy Key and Bestseller Lists
5. Accessing Pogil Bond Energy Key Free and Paid eBooks

- Pogil Bond Energy Key Public Domain eBooks
- Pogil Bond Energy Key eBook Subscription Services
- Pogil Bond Energy Key Budget-Friendly Options
- 6. Navigating Pogil Bond Energy Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Pogil Bond Energy Key Compatibility with Devices
  - Pogil Bond Energy Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pogil Bond Energy Key
  - Highlighting and Note-Taking Pogil Bond Energy Key
  - Interactive Elements Pogil Bond Energy Key
- 8. Staying Engaged with Pogil Bond Energy Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pogil Bond Energy Key
- 9. Balancing eBooks and Physical Books Pogil Bond Energy Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pogil Bond Energy Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Bond Energy Key
  - Setting Reading Goals Pogil Bond Energy Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Bond Energy Key
  - Fact-Checking eBook Content of Pogil Bond Energy Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### **14. Embracing eBook Trends**

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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