#### Buffers

How can a solution neutralize both acids and bases?

#### Why?

Buffer solutions are a mixture of substances that have a fairly constant pH regardless of addition of acid or base. They are used in medicine, industry and manufacturing to keep the pH of a solution stable regardless of other reactions that might be occurring. You have buffers in your blood, for example, because the proper function of enzymes is dependent on the pH of your blood staying in a narrow range. Eating, breathing and exercising can all alter the acidity or basicity of your blood, but the buffers in your blood help control the pH and keep all of your enzymes working properly.

# Model 1 - Comparing Solutions

Beaker A: 100 mL Pure Water				Beaker B: 100 mL of 1.00 M HCl			
H <sub>2</sub> O +	H <sub>2</sub> O ↔ H <sub>3</sub> O	O1*(aq) + OH	1-(aq)	HCl(aq) +	$H_2O \rightarrow$	H <sub>3</sub> O <sup>1+</sup> (aq) +	Cl <sup>1-</sup> (aq)
W54	WB	C4 C	CB .	SA	WB	C4	CB
	1.00	×10 <sup>-7</sup> M 1.00×	10~ <sup>7</sup> M	~0 M		1.00 M	1.00 M
Beaker C: 100 mL of 1.00 M HNO <sub>2</sub> $K_* = 4.6 \times 10^{-4}$				Beaker D: 100 mL of 1.00 M HNO <sub>2</sub> with 100 mL of 1.00 M NaNO <sub>2</sub>			
HNO <sub>2</sub> (	aq) + H <sub>2</sub> O «	→ H <sub>5</sub> O <sup>1+</sup> (aq)	+ NO <sub>2</sub> 1-(aq)	HNO <sub>2</sub> (aq)	+ H <sub>2</sub> O •	→ H <sub>5</sub> O <sup>1+</sup> (aq)	+ NO <sub>2</sub> 1-(aq)
W/4	WB	CA	CB	WM	WB	CA	CB
0.98 A	ď	0.021 M	0.021 M	0.50 M		4.6×10 <sup>→</sup> M	0.50 M

- For each beaker in Model 1 write SA (strong acid), SB (strong base), WA (weak acid), WB (weak base), CA (conjugate acid) or CB (conjugate base) under each molecule or ion in the reaction. See Model 1.
- For each beaker in Model 1 write the final or equilibrium concentrations of the molecules or ions under each species in the reaction. Divide the work among group members. If calculations are required, show your work below.

See Model 1.

Imagine that 1.00 M sodium hydroxide was added dropwise to each of the beakers in Model 1.
List the species in each beaker that would react with this base and neutralize significant quantities (at least 0.05 mole) of it. If neutralization is not likely to occur, write "none."

Beaker A	Beaker B	Beaker C	Beaker D
None	$H_3O^{1*}$	$HNO_2$	$HNO_2$

# **Pogil Buffers Answer Key Chemistry**

Zayika Wilson-Kennedy, Goldie S. Byrd, Eugene Kennedy, Henry T. Frierson

#### **Pogil Buffers Answer Key Chemistry:**

Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more \*\*Broadening Participation in STEM\*\* Zayika Wilson-Kennedy, Goldie S. Byrd, Eugene Kennedy, Henry T. Frierson, 2019-02-28 This book reports on high impact educational practices and programs that have been demonstrated to be effective at broadening the participation of underrepresented groups in the STEM disciplines

The book delves into Pogil Buffers Answer Key Chemistry. Pogil Buffers Answer Key Chemistry is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and indepth insights into Pogil Buffers Answer Key Chemistry, encompassing both the fundamentals and more intricate discussions.

- 1. This book is structured into several chapters, namely:
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  - Chapter 3: Pogil Buffers Answer Key Chemistry in Everyday Life
  - Chapter 4: Pogil Buffers Answer Key Chemistry in Specific Contexts
  - ∘ Chapter 5: Conclusion
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- 6. In chapter 5, this book will draw a conclusion about Pogil Buffers Answer Key Chemistry. This chapter will summarize the key points that have been discussed throughout the book.
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