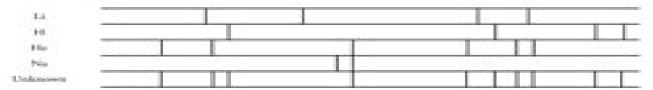


- 24. The hydrogen spectral lines in Model 2 are only the wavelengths of light that are in the visible range and therefore "seen" by the naked eye. However, many other wavelengths can be detected with special equipment.
 - a. Propose a hydrogen electron transition that involves light with a wavelength in the ultraviolet (UV) range (10–400 nm). Any electron transition the ends up on net will edmiran WV wavelength.
 - 6. Propose a hydrogen electron transition that involves light with a wavelength in the infrared (IR) range (1000–106 nm). Anny electron transition, the ends up on Prop not not sell edmit on IR wavelength.
- Below are diagrams for the bright line spectra of four elements and the spectrum of a mixture of unknown pass.



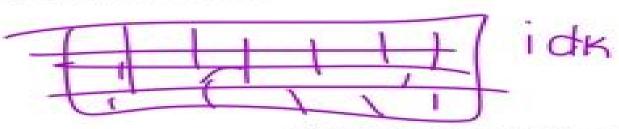
a. Which element(s) are not present in the Unknown!

Li * Na

is. Which element(s) are in the Unknown?

He + H

26. Model 2 shows the emission spectra for hydrogen and boron. Scientists can also record the absorption spectra for elements. Propose how this might be done, and what the absorption spectra of hydrogen and boron would look like.



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Marilyn E. Usher

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