

## Worksheet – Calorimetry

Calorimetry is the experimental measurement of heat ( $q$ ) produced in chemical and physical processes. Heat can not be measured directly, but temperature changes can be measured. The factor that links these two is **heat capacity**.

Heat capacity,  $C$ , is defined as the heat required to raise the temperature of a substance one degree. It relates a measured change in temperature ( $\Delta T$ ) to heat flow ( $q$ ) by the equation:

$$q = C \times \Delta T \times \text{mass}$$

This can be calculated on a per gram basis (specific heat capacity) or a per mole basis (molar heat capacity).

It is easiest to understand the molecular basis for heat capacity by using molar heat capacities of different substances. As energy enters a system, it is distributed among all of the degrees of freedom within a molecule (translational, rotational and vibrational modes). The more complex the molecule, the larger its heat capacity, and the smaller the rise in temperature, since the energy is spread out among more modes.

1. 10 gram samples of iron at a temperature of  $100^\circ\text{C}$  are placed into three beakers each containing 1.0 mol of three different solvents, octane, trichloromethane and methanol, all at  $25.0^\circ\text{C}$ . Their structures and molar masses are shown below.



octane  $\text{C}_8\text{H}_{18}$   
( $\text{mw} = 114 \text{ g/mol}$ )



trichloromethane  $\text{CHCl}_3$   
( $\text{mw} = 119 \text{ g/mol}$ )



methanol  $\text{CH}_3\text{OH}$   
( $\text{mw} = 32 \text{ g/mol}$ )

- How many degrees of freedom in each compound?  $3 \times n$   $n = \# \text{ atoms}$   
octane  $26 \times 3 = 78$  trichloromethane  $3 \times 5 = 15$  methanol  $3 \times 6 = 18$
- Which will have the largest molar heat capacity? *most modes - octane*
- This compound have the (largest, smallest) change in temperature.
- Which will have the largest change in temperature? *smallest  $C$  = trichloro methane*
- The temperature of the methanol rises to  $28.9^\circ\text{C}$ . The specific heat capacity of iron is  $0.449 \text{ J/g}^\circ\text{C}$ . What are the **specific and molar heat capacities** of methanol?  
 $q_{\text{metal}} = -q_{\text{solvent}}$   $1 \text{ mol} = 32 \text{ g}$

$$\begin{aligned}
 (0.449 \text{ J/g}^\circ\text{C})(10 \text{ g})(28.9 - 100) &= -(C)(32 \text{ g})(28.9 - 25) \\
 C &= \frac{2.56 \text{ J}}{\text{g}^\circ\text{C}} \times \frac{32 \text{ g}}{\text{mol}} = 81.9 \text{ J/mol}^\circ\text{C}
 \end{aligned}$$

# Pogil Heat Temperature Calorimetry Answer Key

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**Heat & Temperature Measurement**, 1982      **Animal and Human Calorimetry** J. A. McLean, G. Tobin, 1987 This volume provides a comprehensive survey of the theory practice and techniques of calorimetry as applied to the study of energy metabolism in humans and animals Calorimetry is used to estimate nutritional requirements of man and farm livestock and to evaluate different foods It is also a powerful tool used in research into fundamental nutritional and physiological life processes and in the evaluation of stresses imposed by abnormal or severe environments It is currently being applied in various branches of medical research and can be used as a diagnostic tool in hospitals for investigation of metabolic disorders The authors discuss both direct calorimetry which measures heat loss directly and indirect calorimetry where heat loss is inferred by measurement of some of the chemical byproducts of metabolism In addition guidance is provided to the instrumentation technical problems and precautions necessary to obtain accurate calorimetric measurements

**The Modern Calorimeter** Walter P. White, 1986      Calorimetry Frank H. Bates, 1902      High-temperature Heat-capacity Measurements and Critical Property Determinations Using a Differential Scanning Calorimeter W. V. Steele, Naval Air Propulsion Test Center, United States. Office of the Assistant Secretary for Fossil Energy, National Institute for Petroleum and Energy Research (Bartlesville, Okla.), IIT Research Institute, 1989      High Temperature Solution Calorimetry Patrick Tze-Sheung Chan, 1977      *Specific Heat Measurements by the Method of Mixtures* Estel Burdell Penrod, 1953      *Measuring Heat and Temperature* Arthur Spencer Callen, August Ulmann, International Correspondence Schools, 1932      **Principles of Thermal Analysis and Calorimetry** Simon Gaisford, Vicky Kett, Peter J. Haines, 2016

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