

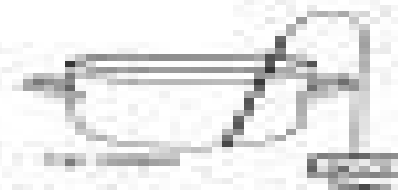
### Calorimetry

What is calorimetry? How does it work? What is its purpose?

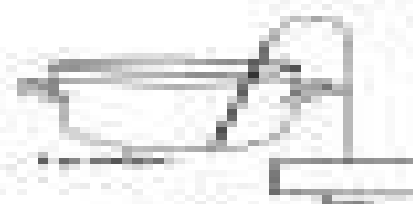
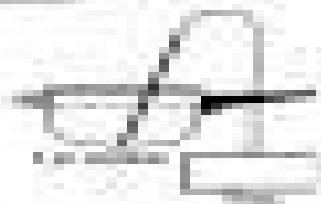
When a substance is heated, the temperature of that substance increases. Will the amount of energy (heat) required to heat different substances to the same temperature increase? Will the same amount of energy be needed to heat different temperatures? Increase or decrease amount of the same substance to different temperatures will require less mass, temperature, heat energy, and the type of substance are related.

#### Model 1 - A Plot of Mass

Calorimetry



The energy (heat) required to heat 100g of water per degree



1. In Model 1, which substance has the most mass of water? 100g

2. Compare the graphs (curves) in Model 1.

a. How many (what units) of energy were added to the substance? 100 J/gram

b. How many (what units) of energy were added to the substance? 100 J

c. In Model 1, which substance has the highest mass? Which substance has the lowest mass? Explain your reasoning.

Why (explain) the above substance? Water, 100g, 100 J/gram, 100 J  
 Explain the above substance? Water, 100g, 100 J/gram, 100 J

# Pogil High School Chemistry Calorimetry Answer Key

**M Tight**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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