

Relative Mass and the Mole

How can atoms be counted using a balance?

Why?

Consider the following equation for a chemical reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

This can be interpreted as two molecules of hydrogen and one molecule of oxygen combining to form two water molecules. But how often do chemists limit their reactions to one or two molecules? Usually a reaction is done with an unimaginable number of molecules. How then do chemists know they have the right mix? The molecules need to be quickly counted! How do we count molecules? The answer is the unit called the **mole**. This activity will start by considering two egg farmers (a chicken farmer and a quail farmer). They produce such large numbers of eggs that they can't count them all individually, so they count in dozens of eggs in some cases, while in other cases they use mass. Weighing is often easier than counting!

Model 1 – Eggs

Chicken		Quail		Ratio of numbers of eggs	Ratio of masses of eggs
Number of eggs in the sample	Mass of the sample	Number of eggs in the sample	Mass of the sample		
1	37.44 g	1	2.34 g	1 : 1	16 : 1
10	374.40 g	10	23.40g	1 : 1	16 : 1
438	16398.72g	438	1024.92g	1 : 1	16 : 1
1 dozen	449.28g	1 dozen	28.08g	1 : 1	16 : 1
1 million	37440000.00g	1 million	2340000.00g	1 : 1	16 : 1

- Consider the data in Model 1.
 - What is the mass of a standard chicken egg?
37.44g
 - What is the mass of a standard quail egg?
2.34g
 - Show mathematically how the 16:1 ratio of masses was calculated in the last column of Model 1.
Ratio of masses of eggs = mass of 1 chicken egg / mass of 1 quail egg
= $37.44\text{g} / 2.34\text{g}$
= $16 / 1$
= $16 : 1$
- Use a calculator to complete the table in Model 1. Divide the work among group members. Reduce all ratios to the lowest whole numbers possible.

Pogil Relative Mass And The Mole Answers

Nigel P. Freestone



Pogil Relative Mass And The Mole Answers:

Counting Moles Nigel P. Freestone, 2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations

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Table of Contents Pogil Relative Mass And The Mole Answers

1. Understanding the eBook Pogil Relative Mass And The Mole Answers
 - The Rise of Digital Reading Pogil Relative Mass And The Mole Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Relative Mass And The Mole Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Relative Mass And The Mole Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Relative Mass And The Mole Answers
 - Personalized Recommendations
 - Pogil Relative Mass And The Mole Answers User Reviews and Ratings
 - Pogil Relative Mass And The Mole Answers and Bestseller Lists
5. Accessing Pogil Relative Mass And The Mole Answers Free and Paid eBooks
 - Pogil Relative Mass And The Mole Answers Public Domain eBooks
 - Pogil Relative Mass And The Mole Answers eBook Subscription Services
 - Pogil Relative Mass And The Mole Answers Budget-Friendly Options
6. Navigating Pogil Relative Mass And The Mole Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Pogil Relative Mass And The Mole Answers Compatibility with Devices
 - Pogil Relative Mass And The Mole Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Relative Mass And The Mole Answers
 - Highlighting and Note-Taking Pogil Relative Mass And The Mole Answers
 - Interactive Elements Pogil Relative Mass And The Mole Answers
 8. Staying Engaged with Pogil Relative Mass And The Mole Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Relative Mass And The Mole Answers
 9. Balancing eBooks and Physical Books Pogil Relative Mass And The Mole Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Relative Mass And The Mole Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Relative Mass And The Mole Answers
 - Setting Reading Goals Pogil Relative Mass And The Mole Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Relative Mass And The Mole Answers
 - Fact-Checking eBook Content of Pogil Relative Mass And The Mole Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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