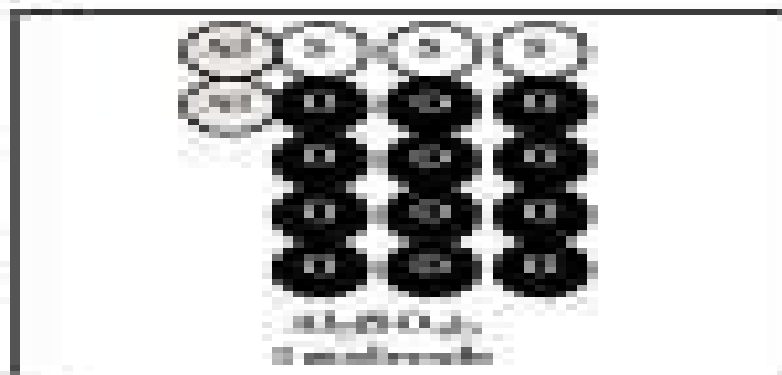


# Balancing Chemical Equations

### Why?

In an unbalanced equation, matter is not conserved (conservation of mass). This means all the reactants (in an unbalanced equation) must be changed up like matter all the products. H<sub>2</sub>O, and possibly the other, unbalanced equations with 2 atoms of hydrogen, need products 2 atoms of each. Similarly, in an unbalanced equation, atoms of oxygen with 2 atoms of hydrogen, need products only 2 atoms of hydrogen. Always an unbalanced equation, is not a chemical equation, and must be balanced before it can be used.

### Model 1:



- 1) According to Model 1, how many atoms of hydrogen and oxygen are in 1 molecule of water,  $\text{H}_2\text{O}$ ?
- 2) According to Model 1, how many atoms of hydrogen and oxygen are in 1 molecule of  $\text{H}_2\text{O}_2$ , hydrogen peroxide?
- 3) According to Model 1, how many atoms of hydrogen, oxygen and oxygen are in 12 molecules of water,  $12 \text{H}_2\text{O}$ ?

# Pogil Activity For Balancing Equations

**Rachel S Tattersall**



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