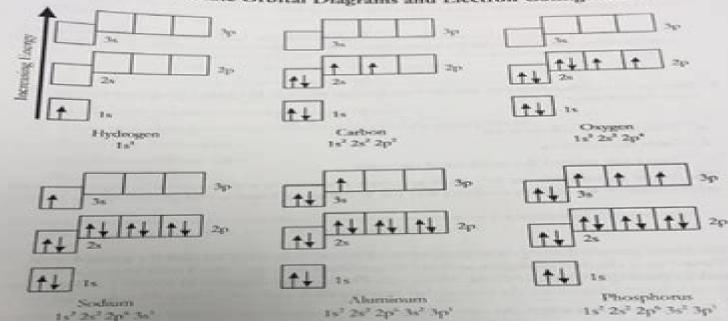
Model 2 - Ground State Orbital Diagrams and Electron Configurations



- Examine the orbital diagrams and electron configurations in Model 2. Match each symbol below with its meaning.

- 1. Single electron
- ___^ †
- II. Pair of electrons with opposite spins
- 1
- III. Atomic orbital (region of space where an electron is likely to be found)
- IV. Sublevel (set of orbitals having equivalent energy)
- r. 1s2 2s2 2p4 V. Electron configuration

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Julius Sumner Miller

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Electrons in Molecules Jean-Pierre Launay, Michel Verdaguer, 2017-12-08 This book provides the reader with a unified understanding of the rapidly expanding field of molecular materials and devices electronic structures and bonding magnetic electrical and photo physical properties and the mastering of electrons in molecular electronics. This revised edition includes updates and additions on hot topics such as molecular spintronics the role of spin in electron transport and molecular machines how electrons can generate molecular motions. Chemists will discover how to understand the relations between electronic structures and properties of molecular entities and assemblies and to design new molecules and materials. Physicists and engineers will realize how the molecular world fits in with their need for systems flexible enough to check

theories or provide original solutions to exciting new scientific and technological challenges. The non specialist will find out how molecules behave in electronics at the most minute sub nanosize level Atomic and Molecular Electronic Configuration Revisited Harold J. Teague, 2014-06-11 In AMEC chapter one you will find a structure for O2 supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O2 fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O2 but also for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the NO and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be2 BeB B2 BC C2 CN N2 and F2 are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O2 OF F2 and F2 F F0 Section IV Photographs of NO and O2 molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N2 going to N2 and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH2 H CH3 H CH4 saturated and tetrahedral arrangement and N H NH H NH2 H NH3 H NH4 also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp2 and sp3 hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added **Introduction to the Electron Theory of Small Molecules** Andrew Crowther Hurley, 1976 **Electron Configuration in the Elements ...** Julius Sumner Miller, 1933

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