Name: Date: Period:

POGIL- Stoichiometry

10 " mosquitoes? He got Mol-aria

What happened to

Avogadro when he got bit by 6.02 x

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called <u>Stoichiometry</u> (stoi-key-ah-meh-tree) Another way of looking at it is using the <u>mole ratio</u> from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A <u>mole ratio</u> is a conversion factor <u>derived from the coefficients of a balanced chemical equation</u> interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: How do I convert between different chemical species in a given reaction?

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole Question: Given the following equation: $1 N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$ How many moles of H_2 are needed to produce 34.8 moles of NH_3 ? Mole ratio Answer: $34.8 \text{ mol NH}_3 \mid 3 \text{ mol H}_2 = 52.2 \text{mol NH}_2$ OR $34.8 \text{ mol NH}_3 \times 3 \text{ mol H}_2 = 52.2 \text{mol NH}_3$

Key Questions:

- I. What was the given in the example above?
- Where did the ^{1 molN₂}/_{2 molNN₃}, come from?
- Double check to see if 52.2 moles of H₂ is the correct answer. <u>SHOW</u> the math below.

Pogil Basic Stoichiometry Key

Emil J. Margolis

Pogil Basic Stoichiometry Key:

Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to covert various quantities to moles or amount of substance This particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in con trasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

Pogil Basic Stoichiometry Key Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the ability of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Pogil Basic Stoichiometry Key**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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