









Model 2 - Proposed Calculations for Mass of NH, to Mass of N,

Toby's Method

$$\frac{\text{x grams}}{50.0 \text{ g}} = \frac{1 \text{ mole N}_2}{2 \text{ moles NH}_4} \rightarrow \text{x = } \underline{-15} \text{ g N}_2$$

Ruchel's Method

$$30.0 \text{ g NH}_3 \times \frac{1 \text{ mole NH}_3}{17.0 \text{ g NH}_4} \times \frac{1 \text{ mole N}_2}{2 \text{ moles NH}_4} \times \frac{28.0 \text{ g N}_2}{1 \text{ mole N}_4} = \frac{28.7 \text{ g N}_2}{1 \text{ mole N}_4}$$

- Model 2 shows three proposed calculations to solve the problem in Question 10. Complete the calculations in Model 2 by filling in the underlined values.
- Which method does not use the mole ratio in an appropriate manner! Explain.
 Total's salted while most breefly to the sales sales while in teament.
- Two of the methods in Model 2 give the same answer. Show that they are mathematically equivalent methods.

 Use either Rachel or Jerry's method from Model 2 to calculate the mass of hydrogen needed to make 30.0 g of ammonia. N_i(g) = 301_i(g) → 2NH_i(g)



Pogil Mole Ratios Model 2

Lifeliqe

Pogil Mole Ratios Model 2:

<u>Mole Ratios</u> Lifeliqe,2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions

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