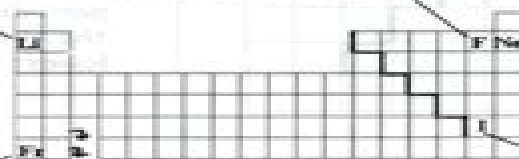


Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

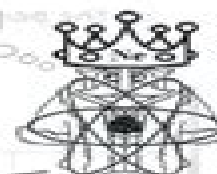
For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.

I am so perfect... I have a full electron shell. I don't need those other elements... I hope they just leave me alone!



Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

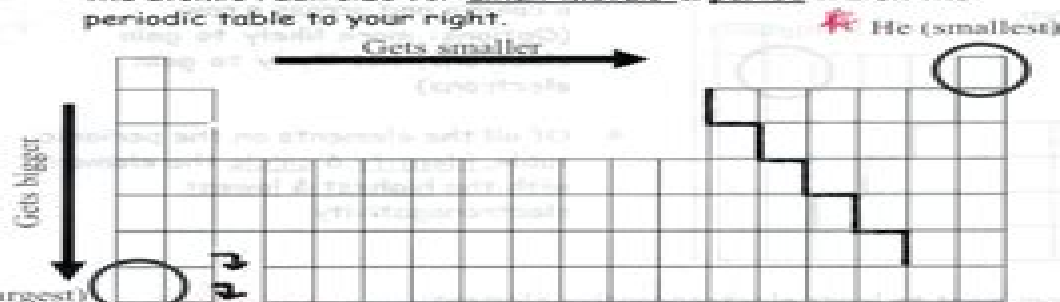
Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.

not a good enough answer
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Table of Contents Pogil Periodic Trends In Atomic Properties Worksheet Answers

1. Understanding the eBook Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - The Rise of Digital Reading Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Personalized Recommendations
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers User Reviews and Ratings
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers and Bestseller Lists
5. Accessing Pogil Periodic Trends In Atomic Properties Worksheet Answers Free and Paid eBooks
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers Public Domain eBooks
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers eBook Subscription Services
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers Budget-Friendly Options

6. Navigating Pogil Periodic Trends In Atomic Properties Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers Compatibility with Devices
 - Pogil Periodic Trends In Atomic Properties Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Highlighting and Note-Taking Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Interactive Elements Pogil Periodic Trends In Atomic Properties Worksheet Answers
8. Staying Engaged with Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Periodic Trends In Atomic Properties Worksheet Answers
9. Balancing eBooks and Physical Books Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Periodic Trends In Atomic Properties Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Setting Reading Goals Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Fact-Checking eBook Content of Pogil Periodic Trends In Atomic Properties Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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