

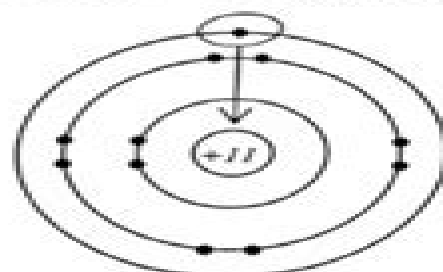
Advanced Periodic Trends

How do the particles in an atom interact to dictate periodic properties of elements?

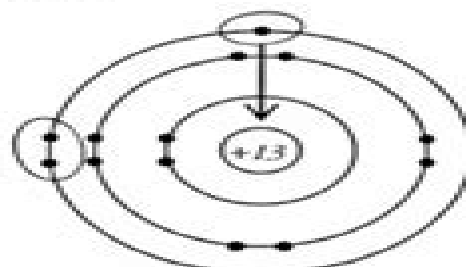
Why?

Previously you have learned about Coulombic attraction and how it governs the trends in properties among the elements of the Periodic Table. For example, as you move across a period, atoms have more protons in the nucleus, which pulls the electrons in tighter making smaller, more electronegative atoms. As you proceed down a column of the periodic table, the valence electrons get further away, and the atoms get larger and less electronegative. Inquisitive students will always ask, "If more protons in the nucleus pull electrons in tighter, then why don't atoms get smaller when going down a column?" That is a good question, and one that you will investigate in this activity.

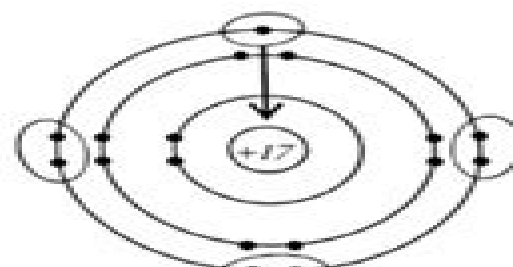
Model 1 – Period 3 Elements



Sodium



Aluminum



Chlorine

Atomic Number	11	13	17
Number of Electrons	11	13	17
Number of Valence Electrons	1	3	7
Core Charge	+1	+3	+7
Atomic Radius	186 pm	143 pm	99 pm
1st Ionization Energy	496 kJ/mol	578 kJ/mol	1251 kJ/mol
Electronegativity	0.9	1.5	3.0

1. Describe the relationship of the three elements in Model 1 with regard to their relative positions on the periodic table.

The three elements are in the same row of the periodic table.

2. Refer to a periodic table to complete the first three rows of the table in Model 1.

See Model 1.

Pogil Periodic Trends Packet

Robert Thomas Sanderson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Pogil Periodic Trends Packet:

Chemistry and Our Universe , Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there *Chemical Periodicity* Robert Thomas Sanderson,2013-04-20 Trends in the Periodic Table Open University. S25- Course Team,1972

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Table of Contents Pogil Periodic Trends Packet

1. Understanding the eBook Pogil Periodic Trends Packet
 - The Rise of Digital Reading Pogil Periodic Trends Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Periodic Trends Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Periodic Trends Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Periodic Trends Packet
 - Personalized Recommendations
 - Pogil Periodic Trends Packet User Reviews and Ratings
 - Pogil Periodic Trends Packet and Bestseller Lists
5. Accessing Pogil Periodic Trends Packet Free and Paid eBooks

- Pogil Periodic Trends Packet Public Domain eBooks
- Pogil Periodic Trends Packet eBook Subscription Services
- Pogil Periodic Trends Packet Budget-Friendly Options
- 6. Navigating Pogil Periodic Trends Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Periodic Trends Packet Compatibility with Devices
 - Pogil Periodic Trends Packet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Periodic Trends Packet
 - Highlighting and Note-Taking Pogil Periodic Trends Packet
 - Interactive Elements Pogil Periodic Trends Packet
- 8. Staying Engaged with Pogil Periodic Trends Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Periodic Trends Packet
- 9. Balancing eBooks and Physical Books Pogil Periodic Trends Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Periodic Trends Packet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Periodic Trends Packet
 - Setting Reading Goals Pogil Periodic Trends Packet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Periodic Trends Packet
 - Fact-Checking eBook Content of Pogil Periodic Trends Packet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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