

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave we need to know how they are bonded, strong
- In order to know a substance behaves like ionic, strong, must be able to separate its ions completely

Electronegativity

- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atom is more able to attract electrons in a bond, electronegative atoms highly electronegative
- Electronegativity increases with increasing atomic number
- From electronegative trends, electronegativity increases with increasing atomic number



Chemical bonding

- Chemical bonding: how substances behave, resulting from types of bonding
- From electronegativity trends

Electronegativity



Electronegativity is a measure of an atom's ability to attract electrons towards itself in a covalent bond. It is a relative property, with fluorine being the most electronegative element and cesium being the least electronegative element.

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