

# Qualitative Analysis and Chemical Bonding

## Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

## Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave we must know how atoms are held together, bonding.
- In order to know a substance behaves how, chemical bonding must be known to explain its behavior.

### Electronegativity

- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent bond.
- More electronegative atoms are more likely to attract electrons in a covalent bond, making them negatively charged ions.
- Electronegativity increases with the strength of a bond.
- From electronegativity trends, electronegativity increases with the strength of a bond.



### Chemical bonding

Chemical bonding: how substances behave, resulting from bonds of electrons, from one element, chemical bonds.

**Electronegativity** is a group of atoms, ions, molecules, and chemical bonds.



Electronegativity is a chemical bond. It is a group of atoms, ions, molecules, and chemical bonds. It is a group of atoms, ions, molecules, and chemical bonds.

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# Qualitative Analysis And Chemical Bonding Review Questions

**Tao Wei**



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