

# Qualitative Analysis and Chemical Bonding

## Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

## Introduction

- Types of atoms are not equal in chemical bond: **electronegativity**
- In order to understand bond, all atoms must be equal in weight, so this factor has been to explain energy
- In order to understand chemical bonding, must use ability to explain the bond formation

### Electronegativity

- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atom is more likely to attract electrons in a bond, making it more negative charge state
- Electronegativity increases with the strength of a bond
- From electronegativity, determine **electronegativity difference**
- From electronegativity difference, determine **chemical bond type**



### Chemical bonding

- Chemical bonding: the attraction forces holding atoms together in a molecule
- From electronegativity difference, determine **chemical bond type**

### Electronegativity



Electronegativity is a measure of an atom's ability to attract electrons towards itself in a covalent bond. It is a relative property, with fluorine being the most electronegative element and cesium being the least electronegative element.

Electronegativity difference is a measure of the difference in electronegativity between two atoms in a covalent bond. It is a relative property, with fluorine being the most electronegative element and cesium being the least electronegative element.

# Qualitative Analysis And Chemical Bonding Answers

**Clyde Metz**



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