

Chemistry - Nail Lab

Purpose:

The purpose is to determine the ratio of copper produced to iron consumed in a replacement reaction.

Procedure:

Day 1

1. Label, then rinse a 250 mL beaker.
2. Put between 6.0 and 6.5 g of copper(II) chloride in the beaker.
3. Add about 80 mL distilled water to the beaker. Stir to dissolve the solid.
4. Mass 2 or 3 nails together to ~ 0.50g.
5. Place the nails in the copper chloride solution. Observe the reactions; record your observations. Place the labeled beaker in the place designated by your teacher.

Day 2

6. Remove the nails. Rinse or scrape the precipitate (copper metal) from the nails into your labeled 250 mL beaker. Place the nails in a labeled small beaker. Note the appearance of the nails. Place this beaker in the drying oven.
7. Decant solution from the 250 mL beaker. Rinse the precipitate with about 25 mL of distilled water. Try to lose as little of the solid copper as you can when you decant. After a 2nd rinse with distilled water, rinse the copper with 25 mL of 1 M HCl. Rinse one last time with distilled water. Then place the labeled beaker in the drying oven.
8. Mass the dry nails, then discuss the nails.

Day 3

9. Mass the beaker + dry copper. Discard the copper in the place designated by your teacher. Wash your beaker and let dry.

Data:

Mass 250 mL beaker	
Mass 250 mL beaker + copper(II) chloride	
Mass nails before reaction	
Mass nails after reaction	
Mass 250 mL beaker + dry copper	

Quantitative Analysis Lab Answers Iron Nails

**South Dakota School of Mines and
Technology**



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