

Give the element and the electron configuration notation for the following:

Quantum Number	Element Identification	Noble Gas Notation
2, 1, +1, $+\frac{1}{2}$	<b>N</b>	<b>[He] 2s<sup>2</sup> 2p<sup>3</sup></b>
3, 2, +1, $-\frac{1}{2}$	<b>Cu</b>	<b>[Ar] 4s<sup>2</sup> 3d<sup>9</sup></b>
5, 3, 0, $-\frac{1}{2}$	<b>Es</b>	<b>[Rn] 7s<sup>2</sup> 5f<sup>11</sup></b>
7, 0, 0, $-\frac{1}{2}$	<b>Ra</b>	<b>[Rn] 7s<sup>2</sup></b>
6, 1, -1, $-\frac{1}{2}$	<b>Po</b>	<b>[Xe] 6s<sup>2</sup> 4f<sup>14</sup> 5d<sup>10</sup> 6p<sup>4</sup></b>
6, 0, 0, $+\frac{1}{2}$	<b>Cs</b>	<b>[Xe] 6s<sup>1</sup></b>
5, 1, +1, $-\frac{1}{2}$	<b>Xe</b>	<b>[Kr] 5s<sup>2</sup> 4d<sup>10</sup> 5p<sup>6</sup></b>
5, 2, -1, $-\frac{1}{2}$	<b>Ir</b>	<b>[Xe] 6s<sup>2</sup> 4f<sup>14</sup> 5d<sup>7</sup></b>
4, 2, -2, $+\frac{1}{2}$	<b>Y</b>	<b>[Kr] 5s<sup>2</sup> 4d<sup>1</sup></b>
1, 0, 0, $-\frac{1}{2}$	<b>He</b>	<b>1s<sup>2</sup></b>

Given the element and quantum number:

Noble-Gas Electron Configuration	Element Identification	Quantum Number
[Rn] 7s <sup>2</sup> 5f <sup>4</sup>	<b>U</b>	<b>5, 3, +1, <math>+\frac{1}{2}</math></b>
[Ar] 4s <sup>2</sup> 3d <sup>10</sup> 4p <sup>2</sup>	<b>Ge</b>	<b>4, 1, 0, <math>+\frac{1}{2}</math></b>
[Ar] 4s <sup>2</sup> 3d <sup>3</sup>	<b>V</b>	<b>3, 2, 0, <math>+\frac{1}{2}</math></b>
[He] 2s <sup>1</sup>	<b>Li</b>	<b>2, 0, 0, <math>+\frac{1}{2}</math></b>
[Kr] 5s <sup>2</sup> 4d <sup>3</sup>	<b>Nb</b>	<b>4, 2, 0, <math>+\frac{1}{2}</math></b>

# Quantum Numbers Practice

**Olga I. Agapova**



## Quantum Numbers Practice:

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