

Collision Theory - Impact for a Chemical Reaction

Why?

The collision theory states that a chemical reaction can only occur between particles when they collide (hit each other). The collision between reactant particles is necessary but not sufficient for a reaction to take place. The collisions also have to be effective. It is important to understand the exact nature of an effective collision since this determines whether or not particles actually react with each other and form new products.

Learning Objectives

- Identify the requirements needed for a successful reaction to occur between reactant particles.

Success Criteria

- Explain the meaning of an effective collision.
- Explain the requirements needed for a reaction to occur between reactant particles.

Resources

- Judith Gould, *Three Strikes Equals a Hit*, STANYS Science Teacher Bulletin (2000).
- Zumdahl, Zumdahl and DeCoste. 2002. *World of Chemistry*. Houghton Mifflin, pp. 537 - 541

Prerequisites

- Chemical reaction nomenclature
- Balancing chemical reactions
- Lewis structures (electron-dot-diagrams)

New Concepts

- Collision theory
- Effective collision
- Activation energy

Pogil Collision Theory

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