

1) State the four types of chemical reactions in today's experiment.

2) Complete the table by filling in the formula for the ionic compound formed by each pair of cations and anions, as shown for the first pair.

Ion	$\text{Na}^+$	$\text{Ca}^{2+}$	$\text{Zn}^{2+}$	$\text{Al}^{3+}$
$\text{OH}^-$	$\text{NaOH}$	$\text{Ca(OH)}_2$	$\text{Zn(OH)}_2$	$\text{Al(OH)}_3$
$\text{Cl}^-$	$\text{NaCl}$	$\text{CaCl}_2$	$\text{ZnCl}_2$	$\text{AlCl}_3$
$\text{SO}_4^{2-}$	$\text{Na}_2\text{SO}_4$	$\text{CaSO}_4$	$\text{ZnSO}_4$	$\text{Al}_2(\text{SO}_4)_3$
$\text{PO}_4^{3-}$	$\text{Na}_3\text{PO}_4$	$\text{Ca}_3(\text{PO}_4)_2$	$\text{Zn}_3(\text{PO}_4)_2$	$\text{AlPO}_4$

3) Complete and balance the following reactions:



4) Use the solubility table below for questions a and b.

Soluble Compounds	Exceptions
Compounds containing alkali metals ions and the ammonium ion	
Nitrates, bicarbonates, and chlorates	
Halides	Halides of $\text{Ag}^+$ , $\text{Hg}_2^{2+}$ , and $\text{Pb}^{2+}$
Sulfates	Sulfates of $\text{Ag}^+$ , $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Hg}_2^{2+}$ , and $\text{Pb}^{2+}$
Insoluble Compounds	Exceptions
Carbonates, phosphates, chromates, and sulfides	Compounds containing alkali metals ions and the ammonium ion
Hydroxides	Compounds containing alkali metals ions and the ammonium ion

a) Select the ionic compound which is insoluble in water.

A)  $\text{NaHCO}_3$  B)  $\text{Cu(NO}_3)_2$  C)  $\text{AgBr}$  D)  $\text{MgCl}_2$  E)  $\text{NH}_4\text{Br}$

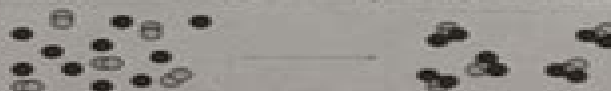
Answer: ( )

b) What is the precipitate that forms when solutions of  $\text{Na}_3\text{PO}_4$  and  $\text{Fe(NO}_3)_3$  are mixed?

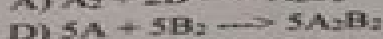
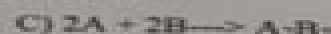
A)  $\text{NaNO}_3$  B)  $\text{NaFe}$  C)  $\text{FePO}_4$  D)  $\text{PO}_4\text{NO}_3$  E)  $\text{Fe}_3\text{PO}_4$

Answer: ( )

5) The reaction of A (dark color spheres) with B (light color spheres) is shown in the following diagram:



Which equation best describes the stoichiometry of the reaction?



Answer: ( )

# Pogil Nuclear Equations Answer Key

**SJ Ball**



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