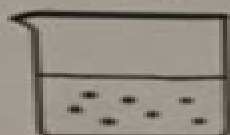


- Which illustration below represents
 - solute particles in a solid state in water?
 - solute particles in an aqueous state?



- What variables are controlled in all five beakers of Model 1?
 - amount of water - volume is the same
 - type of beakers
 - stirred for 2 hours
- Count the particles present in each beaker of Model 1. Fill in the table to show the number of dissolved solute particles and the number of solid solute particles.
- Consider the beakers in Model 1.
 - Which beakers represent unsaturated solutions?
B, D
 - Which beakers represent saturated solutions?
A, C, E
- Beakers A-E in Model 1 are depicted as representing five different or separate solutions. They could also be considered as five "snapshots" of the same beaker over time. In other words, if additional measured quantities of solute were stirred into beaker A in small increments over time, then beakers B-E would result.
 - When a small amount of additional solute is added to an unsaturated solution, what happens to the number of dissolved particles? Provide specific evidence from Model 1 to support your answer.
the number of dissolved particles increases
 - When a small amount of additional solute is added to a saturated solution, what happens to the number of dissolved particles? Provide specific evidence from Model 1 to support your answer.
the dissolved particles stay the same and the solid particles increase
 - Predict what would happen if a small amount of additional solute were stirred into beaker E in Model 1.
more solid particles would stay at the bottom

Pogil Solution Manual Equilibrium

Sebastian Brünink



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Chemical Equilibrium William Guenther, 1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation The Determination of Stability Constants Francis J. C. Rossotti, 1964 Solutions Manual for Robert A. Alberty Physical Chemistry Robert A. Alberty, 1983 Unified Equilibrium Calculations William Benton Guenther, 1991-08-19 A new approach to the academic treatment of solution equilibria is presented The author unifies homonuclear equilibrium calculations in one concept The alpha species fraction and bound proton and bound ligand ratio as a function of a single master variable the unbound H or L yield complete balances A single logic is maintained for all cases by equating the chemical binding expressed as an equilibrium condition and as a material balance condition *Solutions Manual* Joseph Topich, 2003-05-01 Critical Evaluation of Equilibrium Constants in Solution, 1980 *Solutions Manual Physical Methods for Chemists* Ngai M. Wong, Russell S. Drago, Numerical Solution of Dynamic Equilibrium Equations Jamal J. Nasir, 1983

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