

Gas Variables

How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 1 atm
Temperature = 200 K



A2

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 2 atm
Temperature = 200 K



A3

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 3 atm
Temperature = 200 K

Experiment B (Heating the gas)



B1

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 1 atm
Temperature = 200 K



B2

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 2 atm
Temperature = 400 K



B3

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 3 atm
Temperature = 600 K

**Note: Volume in this model is recorded in units rather than liters because 4 molecules of gas at the conditions given would occupy a very small space ($\sim 1 \times 10^{-22}$ mL). The particles shown here are much larger compared to the space between them than actual gas particles.*

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Table of Contents Pogil Chemistry Answer Key For Gas Variables

1. Understanding the eBook Pogil Chemistry Answer Key For Gas Variables
 - The Rise of Digital Reading Pogil Chemistry Answer Key For Gas Variables
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Chemistry Answer Key For Gas Variables
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Chemistry Answer Key For Gas Variables
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Chemistry Answer Key For Gas Variables
 - Personalized Recommendations

- Pogil Chemistry Answer Key For Gas Variables User Reviews and Ratings
- Pogil Chemistry Answer Key For Gas Variables and Bestseller Lists
- 5. Accessing Pogil Chemistry Answer Key For Gas Variables Free and Paid eBooks
 - Pogil Chemistry Answer Key For Gas Variables Public Domain eBooks
 - Pogil Chemistry Answer Key For Gas Variables eBook Subscription Services
 - Pogil Chemistry Answer Key For Gas Variables Budget-Friendly Options
- 6. Navigating Pogil Chemistry Answer Key For Gas Variables eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Chemistry Answer Key For Gas Variables Compatibility with Devices
 - Pogil Chemistry Answer Key For Gas Variables Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Chemistry Answer Key For Gas Variables
 - Highlighting and Note-Taking Pogil Chemistry Answer Key For Gas Variables
 - Interactive Elements Pogil Chemistry Answer Key For Gas Variables
- 8. Staying Engaged with Pogil Chemistry Answer Key For Gas Variables
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Chemistry Answer Key For Gas Variables
- 9. Balancing eBooks and Physical Books Pogil Chemistry Answer Key For Gas Variables
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Chemistry Answer Key For Gas Variables
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Chemistry Answer Key For Gas Variables
 - Setting Reading Goals Pogil Chemistry Answer Key For Gas Variables
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Chemistry Answer Key For Gas Variables
 - Fact-Checking eBook Content of Pogil Chemistry Answer Key For Gas Variables

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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