

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 Broken	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 966.7 Formed
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 Broken	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 Formed
C)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 Broken	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 1,470.4 Formed
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 Broken	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 Formed

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify. **A, B, C, D**
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify. **E, F, G, H**
- Circle the correct word to complete each sentence below.
 - When bonds are (**broken**/formed) there is a positive energy change.
 - Breaking bonds is (**endothermic**/exothermic).
 - When bonds are (broken/**formed**) there is a negative energy change.
 - Forming bonds is (endothermic/**exothermic**).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
 - How do the changes in energy for the reverse reactions compare?
When the reaction is reversed the **sign** is reversed but the change in energy value remains the same.
 - Explain your answer to part *a* considering what you learned from Questions 1–3 about bond breaking and bond formation.
The breaking of bonds requires energy to be absorbed and forming bonds involves the release of energy.



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Riley TARA



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Table of Contents Pogil Extension Questions Bond

1. Understanding the eBook Pogil Extension Questions Bond
 - The Rise of Digital Reading Pogil Extension Questions Bond
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Extension Questions Bond
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Extension Questions Bond
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Extension Questions Bond
 - Personalized Recommendations
 - Pogil Extension Questions Bond User Reviews and Ratings
 - Pogil Extension Questions Bond and Bestseller Lists
5. Accessing Pogil Extension Questions Bond Free and Paid eBooks

- Pogil Extension Questions Bond Public Domain eBooks
 - Pogil Extension Questions Bond eBook Subscription Services
 - Pogil Extension Questions Bond Budget-Friendly Options
6. Navigating Pogil Extension Questions Bond eBook Formats
- ePub, PDF, MOBI, and More
 - Pogil Extension Questions Bond Compatibility with Devices
 - Pogil Extension Questions Bond Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Pogil Extension Questions Bond
 - Highlighting and Note-Taking Pogil Extension Questions Bond
 - Interactive Elements Pogil Extension Questions Bond
8. Staying Engaged with Pogil Extension Questions Bond
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Extension Questions Bond
9. Balancing eBooks and Physical Books Pogil Extension Questions Bond
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Extension Questions Bond
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Extension Questions Bond
- Setting Reading Goals Pogil Extension Questions Bond
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Extension Questions Bond
- Fact-Checking eBook Content of Pogil Extension Questions Bond
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

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- Interactive and Gamified eBooks

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