

Model 2 – Ground State Orbital Diagrams and Electron Configurations



Hydrogen
 $1s^1$



Carbon
 $1s^2 2s^2 2p^2$



Oxygen
 $1s^2 2s^2 2p^4$



Sodium
 $1s^2 2s^2 2p^6 3s^1$



Aluminium
 $1s^2 2s^2 2p^6 3s^2 3p^1$



Phosphorus
 $1s^2 2s^2 2p^6 3s^2 3p^3$

6. Examine the orbital diagrams and electron configurations in Model 2. Match each symbol below with its meaning.

___ a.

___ b.

___ c.

___ d.

___ e. $1s^2 2s^2 2p^4$

I. Single electron

II. Pair of electrons with opposite spins

III. Atomic orbital (region of space where an electron is likely to be found)

IV. Sublevel (set of orbitals having equivalent energy)

V. Electron configuration

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Primrose Kitten

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape with a gradient, partially overlapping the bar's end.

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Electrons in Molecules Jean-Pierre Launay, Michel Verdaguer, 2017-12-08 This book provides the reader with a unified understanding of the rapidly expanding field of molecular materials and devices electronic structures and bonding magnetic electrical and photo physical properties and the mastering of electrons in molecular electronics This revised edition includes updates and additions on hot topics such as molecular spintronics the role of spin in electron transport and molecular machines how electrons can generate molecular motions Chemists will discover how to understand the relations between electronic structures and properties of molecular entities and assemblies and to design new molecules and materials Physicists and engineers will realize how the molecular world fits in with their need for systems flexible enough to check theories or provide original solutions to exciting new scientific and technological challenges The non specialist will find out how molecules behave in electronics at the most minute sub nanosize level

Atomic and Molecular Electronic Configuration Revisited Harold J. Teague, 2014-06-11 In AMEC chapter one you will find a structure for O₂ supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O₂ fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O₂ but also

for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the N O and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be₂ BeB B₂ BC C₂ CN N₂ and F₂ are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O₂ OF F₂ and F₂ F F₀ Section IV Photographs of NO and O₂ molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N₂ going to N₂ and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH₂ H CH₃ H CH₄ saturated and tetrahedral arrangement and N H NH H NH₂ H NH₃ H NH₄ also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp² and sp³ hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added

Electronic Configuration: A Formula Handbook N.B. Singh, Electronic Configuration A Formula Handbook is a concise and indispensable guide for understanding the arrangement of electrons in atoms and molecules This handbook provides clear and easy to follow formulas and rules for determining electronic configurations enabling readers to quickly and accurately predict the distribution of electrons in various atomic and molecular systems Whether you re a student studying chemistry or a professional in the field this book serves as a valuable reference for mastering electronic configurations and their implications in chemical bonding and reactivity

Structures and Approximations for Electrons in Molecules David B. Cook, 1978

Chemistry Uniy 1 Archie Gibb, Arthur A. Sandison, Learning and Teaching Scotland, Andrew Watson, 2000

Atoms, Electrons, Structure and Bonding Primrose Kitten, 2017-06-19 I m constantly telling you the best way to learn is by practicing questions so I ve made you a book full of practice questions Multiple choice questions to reflect the style of exam questions activities to complete equations for you to balance compounds for you to work out the formula for lots of things that you need to recall and practice long answer exam style questions This book is not designed as a text book or revision guide but as a workbook There are lots of good and bad expensive and free revision guides out there on my YouTube channel and other great websites So there is no point in me adding to the masses All the teaching all the new content is available for free on my YouTube channel this book is for you to practice and learn The best way to approach this is to watch the teaching video and make notes or after class try a section and check the answers Any corrections that are needed after the book is published will be listed on my website www.primrosekitten.com these will be corrected in the next version of the

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Introduction to the Electron Theory of Small Molecules Andrew Crowther Hurley, 1976

Chemistry David John Cole-Hamilton, Iain Patterson, 2001

Electron Configuration , Electron Configuration as the Basis of the Periodic Table William Fay Luder, 1943

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