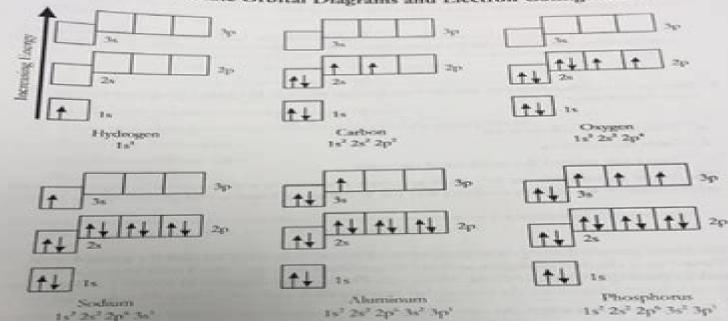
# Model 2 - Ground State Orbital Diagrams and Electron Configurations



- Examine the orbital diagrams and electron configurations in Model 2. Match each symbol below with its meaning.

- 1. Single electron
- \_\_\_^ †
- II. Pair of electrons with opposite spins
- 1
- III. Atomic orbital (region of space where an electron is likely to be found)
- IV. Sublevel (set of orbitals having equivalent energy)
- r. 1s2 2s2 2p4 V. Electron configuration

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for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the NO and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be2 BeB B2 BC C2 CN N2 and F2 are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O2 OF F2 and F2 F F0 Section IV Photographs of NO and O2 molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N2 going to N2 and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH2 H CH3 H CH4 saturated and tetrahedral arrangement and N H NH H NH2 H NH3 H NH4 also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp2 and sp3 hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added **Electronic Configuration: A Formula** Handbook N.B. Singh, Electronic Configuration A Formula Handbook is a concise and indispensable guide for understanding the arrangement of electrons in atoms and molecules This handbook provides clear and easy to follow formulas and rules for determining electronic configurations enabling readers to quickly and accurately predict the distribution of electrons in various atomic and molecular systems Whether you re a student studying chemistry or a professional in the field this book serves as a valuable reference for mastering electronic configurations and their implications in chemical bonding and reactivity Structures and Approximations for Electrons in Molecules David B. Cook, 1978 Chemistry Univ 1 Archie Gibb, Arthur A. Sandison, Learning and Teaching Scotland, Andrew Watson, 2000

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