

## Naming Acids

What makes an acid unique and how are acids named?

### Why?

A variety of acids are used in foods, industry, and research. Acids are covalently bonded molecules, but when they are put into water they produce ions. One of the ions produced is always  $H^+$ , which immediately combines with a water molecule to form the hydronium ion ( $H_3O^+$ ). The  $H_3O^+$  ion is what defines the acidic properties of a substance. Because of their special classification, acids have a naming system different from ionic or other molecular (covalent) compounds.

### Model 1 – Binary Acids

Acid	Name of acid in aqueous solution	Cation (+)*	Anion (–)
HCl	Hydrochloric acid	$H_3O^+$	$Cl^-$
HBr	Hydrobromic acid	$H_3O^+$	$Br^-$
$H_2S$	Hydrosulfuric acid	$2H_3O^+$	$S^{2-}$
HF	Hydrofluoric acid	$H_3O^+$	$F^-$

\*Hydrogen ions ( $H^+$ ) join with water molecules in solution to form hydronium ions,  $H_3O^+$ .

1. Complete the cation and anion columns of the table in Model 1. Be careful to show the charges on the ions.

2. Why does hydrosulfuric acid contain two hydrogens?

*Two hydrogens with a +1 charge are required to balance the 2– charge of the sulfide anion.*

3. Look at the formulas and names of the binary acids in Model 1.

- a. What prefix is used at the beginning of the name for all binary acids?

*Binary acids begin with the prefix “hydro-.”*

- b. What suffix is used at the end of the name for all binary acids?

*The names of binary acids end with “-ic,” followed by the word “acid.”*

4. The prefix “bi” means “two.” Propose a reason that the acids in Model 1 are all referred to as “binary” acids.

*All binary acids have two different elements or atoms in them.*



5. Write a rule for naming binary acids.

*All binary acids begin with the prefix hydro-, which precedes the name of any nonmetal other than hydrogen, in the compound. The name of this nonmetal is modified to end with the suffix -ic.*

*Example: hydro\_\_\_\_\_ic acid.*



### Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (–)	Polyatomic Anion Name
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# Pogil Naming Acids

**Siddappa N.Byrareddy**



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## **Table of Contents Pogil Naming Acids**

1. Understanding the eBook Pogil Naming Acids
  - The Rise of Digital Reading Pogil Naming Acids
  - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Naming Acids
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pogil Naming Acids
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Naming Acids
  - Personalized Recommendations
  - Pogil Naming Acids User Reviews and Ratings
  - Pogil Naming Acids and Bestseller Lists
5. Accessing Pogil Naming Acids Free and Paid eBooks
  - Pogil Naming Acids Public Domain eBooks

- Pogil Naming Acids eBook Subscription Services
- Pogil Naming Acids Budget-Friendly Options
- 6. Navigating Pogil Naming Acids eBook Formats
  - ePub, PDF, MOBI, and More
  - Pogil Naming Acids Compatibility with Devices
  - Pogil Naming Acids Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pogil Naming Acids
  - Highlighting and Note-Taking Pogil Naming Acids
  - Interactive Elements Pogil Naming Acids
- 8. Staying Engaged with Pogil Naming Acids
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pogil Naming Acids
- 9. Balancing eBooks and Physical Books Pogil Naming Acids
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pogil Naming Acids
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Naming Acids
  - Setting Reading Goals Pogil Naming Acids
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Naming Acids
  - Fact-Checking eBook Content of Pogil Naming Acids
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

## 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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