

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Crystal & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave in regard to their chemical bonds, we must first understand the nature of the atoms that form the bonds.
- In order to understand how all substances behave in regard to their chemical bonds, we must first understand the nature of the atoms that form the bonds.

Electronegativity

- Electronegativity is the ability of an atom to attract electrons towards itself in a chemical bond.
- The more electronegative an atom is, the more it will attract electrons towards itself in a chemical bond.
- Electronegativity is a relative property; it is not an absolute value.
- The electronegativity of an atom is determined by its position in the periodic table.



Chemical bonding

- Chemical bonding is the process by which atoms are held together in a molecule or crystal lattice.
- The type of chemical bond that forms between two atoms depends on the electronegativity difference between them.

Electronegativity

- Electronegativity is a measure of an atom's ability to attract electrons towards itself in a chemical bond.



Electronegativity is a measure of an atom's ability to attract electrons towards itself in a chemical bond. The electronegativity difference between two atoms determines the type of chemical bond that forms between them.

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Qualitative Analysis Chemical Bonding Teachers Notes

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Qualitative Analysis Chemical Bonding Teachers Notes:

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