

Chapter 21: Temperature, Heat and Expansion

All matter – solid, liquid and gas – is made of atoms or molecules, which are continually jiggling. As this jiggling is a movement, all these particles must have a **kinetic energy**.

This kinetic energy of all the particles of an object is our concept of **temperature**. When the particles of a substance have a high kinetic energy, they are all jiggling and bouncing around very fast. To us, the substance would feel hot.(?)

Temperature

On the Celsius scale, the freezing and boiling points of water were chosen to be 0° and 100° respectively. This is an arbitrary system used to give us a guide to the temperature of something.

Consider a substance whose particles have no kinetic energy. Due to our definition of temperature, this would mean the substance couldn't be any colder. This temperature is **-273°C** and called **absolute zero**.

Temperature Heat And Expansion Answers

Miguel F. Acevedo



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