

## Stoichiometry

This lab shows how stoichiometry is used in predicting the amount of product in a chemical reaction. The lab is performed for you in the video linked below:

<https://www.youtube.com/watch?v=ATs7C9gIKXg>

The details in the video will help you fill in the data table and answer the questions in this lab. Please have a pen/paper ready when pertinent data or information comes up as you are answering the lab. I have copied down the procedure for you, just in case it might be helpful. Pages 3 and 4 are where you start answering the questions from the data presented in the video.

As the calculations are being demonstrated, please note how this connects to what we covered in our stoichiometry lecture.

# Target Stoichiometry Lab

## Procedure

1. Obtain a large test tube and weigh it on one of the balances. Record this mass in the Data Table.
2. Go back to your lab station and place one large scoop of Baking soda (sodium bicarbonate,  $\text{NaHCO}_3$ ) into the test tube.
3. Using the same scale as before, weigh the test tube with the baking soda. Record this mass in the Data Table.
4. Calculate the mass of the baking soda in the test tube and record the amount in the Data Table.
5. Holding the test tube nearly horizontal, shake the baking soda gently so that it spreads out a bit (see below).



6. Tighten the test tube clamp securely around the test tube, just below the lip so that it is positioned nearly horizontally about 20 cm above the lab desk (see Figure 2).



**Figure 2**

# Target Stoichiometry Answered Lab Sheet

**Rachel Sandford**



## **Target Stoichiometry Answered Lab Sheet:**

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