

Name: Answer Key

Class:

Date:

Identify which atomic model is shown in each picture and write the name below the picture. Write the scientist responsible for the model below each picture as well. There is a word bank below to help you.

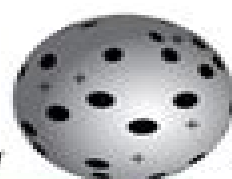
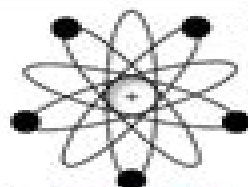
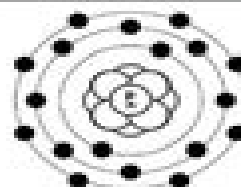
History of the Atom



Indivisible atom

1a. Model: simple, first atomic theoryb. Scientist(s) responsible: John Dalton

Scattered
negatively
charged,
small
particles
within the
blob

positively
charged,
low density,
spacious
blob2a. Model: plum pudding modelb. Scientist(s) responsible: J. J. Thomsonpositively
charged center
with orbiting
negatively
charged particles3a. Model: planetary modelb. Scientist(s) responsible: Ernest RutherfordNucleus with protons,
neutrons, and electrons
that orbit at defined
energy levels and
distances away from
the nucleus4a. Model: Bohr modelb. Scientist(s) responsible: Niels BohrNucleus with electrons
that have calculated
probable areas where
they are located, but no
exact location can be
determined.5a. Model: electron cloud modelb. Scientist(s) responsible: Modern Scientists

6. Now, examine the following statements from Dalton's Atomic Theory. Circle true or false for each statement and if it's false, write below what we know about the atom today that contradicts his theory.

A. All matter is made of atoms.
True or False

B. Atoms are indivisible and indestructible.
True or False

Atoms can be destroyed and sometimes decay.

C. All atoms of a given element are identical in mass and properties.
True or False

Isotopes are atoms of a given element with different masses.

D. Compounds are formed by a combination of two or more different kinds of atoms.
True or False

Models (not in chronological order)

planetary model
simple, first atomic theory
plum pudding model
Bohr model
Electron Cloud model

Scientists (not in order)

Ernest Rutherford
J. J. Thomson

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