

# Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

## PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 2 Hydrogen gas ( $\text{H}_2$ ) reacts with oxygen gas ( $\text{O}_2$ ) to produce water ( $\text{H}_2\text{O}$ ). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3 8.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 4 1.8 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 3.5 grams. Calculate the percent yield of zinc sulfide.

- 5 The actual yield was 28.8 grams. Calculate the percent yield.

- 6 Magnesium ( $\text{Mg}$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce magnesium chloride ( $\text{MgCl}_2$ ) and hydrogen gas ( $\text{H}_2$ ). 10.0 grams of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 7 1.0 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 3.5 grams. Calculate the percent yield of phosphorus trichloride.

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 9 Calcium carbonate ( $\text{CaCO}_3$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce calcium chloride ( $\text{CaCl}_2$ ), carbon dioxide ( $\text{CO}_2$ ), and water ( $\text{H}_2\text{O}$ ). 15.0 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 13.9 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

$$\begin{array}{r} 15.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{100.09 \text{ g CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 13.9 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

- 10 The actual yield was 13.9 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{13.9 \text{ g}}{13.9 \text{ g}} \times 100 = 100\% \text{ CaCl}_2$$

- 11 Potassium ( $\text{K}$ ) reacts with iodine ( $\text{I}_2$ ) to produce potassium iodide ( $\text{KI}$ ). 10.0 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 10.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 166.10 \text{ g KI} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

YIELD

ANSWER KEY INCLUDED

# Percent Yield Example Answer Key

**National Institute of Standards and  
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### **Percent Yield Example Answer Key:**

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*Climate-resilience policies and investments for Egypt's agriculture sector: Sustaining productivity and food security* Perez, Nicostrato D., Kassim, Yumna, Ringler, Claudia, Thomas, Timothy S., ElDidi, Hagar, Breisinger, Clemens, 2021-09-24 The importance of a resilient agriculture sector in providing food security livelihoods and household income was highlighted in many countries by the recent pandemic as was the capacity of the sector to cushion the negative impacts of the subsequent economic slowdown This has been the case in Egypt where agriculture has been resilient to the health crisis in comparison with the service and industry sectors Breisinger et al 2020 However the sector s resiliency is gradually being corroded by climate change with lasting harmful effects for agriculture and food systems

**Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and

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*and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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