

Periodic Trends

Can the properties of an element be predicted using a periodic table?

Why?

The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

3 1. Consider the data in Model 1 on the following page.

a. Each element has three numbers listed under it. Which value represents the atomic radius?

The first number below the diagram

b. What are the units for the atomic radius?

pm picometers 10^{-12} m

c. Write a complete sentence to convey your understanding of atomic radius. *Note:* You may not use the word "radius" in your definition.

Distance from center of atom (nucleus) to outer electrons

2 2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.

Increases $H < Li < Na < K$
 $C < Si < Ge$

$F < Cl < Br$



2 3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. *Hint:* You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.

As the number of electrons increases, new energy levels need to be added further from the nucleus, increasing the radius.

2 4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.

Decreases $Li > Be > B > C > N > O > F > Ne$
 $Na > Mg > Al > Si > P > S > Cl > Ar$



2 5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.

Highest energy level is constant, but the number of protons in the nucleus increases, increasing the charge of the nucleus and the attractive force on the electrons, pulling the highest energy level closer to the nucleus.

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