

AP® CHEMISTRY 2005 EXAMINATIONS CHEMISTRY SECTION

Question 10
(8 points)

As sample of $\text{C}_2\text{H}_2\text{Br}_4$ is subjected to continuous infrared light of 2.5 micrometers, continuous heat is produced. The product of this sample is subsequently subjected to gas chromatography.



The structure $\text{C}_2\text{H}_2\text{Br}_4$ is considered for decomposition at temperatures within 100°C to 200°C . Based on the graph, what is the most likely product of this decomposition?

Answer the two:

- Write the decomposition reaction. Be sure to include the products in parentheses. Use molecular formulas. Indicate physical states.
- Write the decomposition equation. Be sure to include energy change when chemical equation is written. Indicate when a rate constant for a reaction is given.

1 point is earned for each correct answer.

$\text{C}_2\text{H}_2\text{Br}_4$ reacts readily with H_2O to produce $\text{C}_2\text{H}_2\text{Br}_2$ and H_2O in a reaction that is first-order in $\text{C}_2\text{H}_2\text{Br}_4$.

$$\text{C}_2\text{H}_2\text{Br}_4 \rightarrow \text{C}_2\text{H}_2\text{Br}_2 + \text{H}_2\text{O} \quad k = 0.001 \text{ s}^{-1} \text{ at } 100^\circ\text{C}$$

At 100°C , $\text{C}_2\text{H}_2\text{Br}_4$ is expected to react with H_2O to form $\text{C}_2\text{H}_2\text{Br}_2$ and H_2O in a reaction that is first-order in $\text{C}_2\text{H}_2\text{Br}_4$. Based on the rate of reaction, what is the most likely product of this reaction?

- Write the decomposition reaction. Be sure to include the products in parentheses. Use molecular formulas. Indicate physical states.
- Write the decomposition equation. Be sure to include energy change when chemical equation is written. Indicate when a rate constant for a reaction is given.

1 point is earned for each correct answer.

Response Question Ap Chemistry 2013

M Woodhall



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