

Predicting Bond Types

TABLE 4.15 Predicting Bond Type from Electronegativity Differences

Molecule		Electronegativity Difference	Bond Type	Reason
H ₂	H—H	2.1 − 2.1 = 0	Nonpolar covalent	Less than 0.4
Cl ₂	Cl—Cl	3.0 − 3.0 = 0	Nonpolar covalent	Less than 0.4
HBr	$\overset{\delta^+}{\text{H}}-\overset{\delta^-}{\text{Br}}$	2.8 − 2.1 = 0.7	Polar covalent	Greater than 0.4, but less than 1.8
HCl	$\overset{\delta^+}{\text{H}}-\overset{\delta^-}{\text{Cl}}$	3.0 − 2.1 = 0.9	Polar covalent	Greater than 0.4, but less than 1.8
NaCl	Na ⁺ Cl [−]	3.0 − 0.9 = 2.1	Ionic	Greater than 1.8
MgO	Mg ²⁺ O ^{2−}	3.5 − 1.2 = 2.3	Ionic	Greater than 1.8

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brain Affected learners require timely and tailored interventions informed and shaped by neurological findings In this ground
breaking text Professor Butterworth explains the latest research in the science of dyscalculia in a clear non technical way
Crucially he shows that dyscalculia is caused by a core deficit in the ability to accurately and swiftly represent the number of
objects in a set an ability that underpins learning arithmetic and clearly differentiates dyscalculia from other forms of early
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