

## Review #5: Chemical Equations and Stoichiometry

- Define limiting factor. Why is it significant?  
The limiting factor is the reactant that runs out first in a chemical reaction. The limiting factor is significant because once it is gone, the reaction stops. It is the amount of the limiting factor that determines how much product can be produced.
- Find the simplest formula for the compound with composition:
  - 38.7% carbon, 9.7% hydrogen and 51.6% oxygen ( $\text{CH}_3\text{O}$ )
  - 82.4% nitrogen and 17.6% hydrogen ( $\text{NH}_3$ )
- A certain compound is 40.0% carbon, 6.7% hydrogen and 53.3% oxygen by weight. One mole of this substance weighs 180 grams. What is the molecular formula of the compound? ( $\text{C}_6\text{H}_{12}\text{O}_6$ )
- Balance these expressions. Type of Reaction
  - $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$  synthesis
  - $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$  combustion
  - $3 \text{ MgO} + 2 \text{ H}_3\text{PO}_4 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + 3 \text{ H}_2\text{O}$  double displacement
  - $2 \text{ PbO}_2 \rightarrow 2 \text{ PbO} + \text{O}_2$  decomposition
  - $\text{SiO}_2 + 4 \text{ HF} \rightarrow \text{SiF}_4 + 2 \text{ H}_2\text{O}$  double displacement
  - $2 \text{ C}_{10}\text{H}_{22} + 31 \text{ O}_2 \rightarrow 20 \text{ CO}_2 + 22 \text{ H}_2\text{O}$  combustion
  - $\text{Mg} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2 + \text{MgSO}_4$  single displacement
  - $\text{Sb}_2\text{S}_3 + 12 \text{ HCl} \rightarrow 2 \text{ H}_3\text{SbCl}_6 + 3 \text{ H}_2\text{S}$  Can't classify
- For the reaction:  $2 \text{ Fe}_2\text{O}_3 + 3 \text{ C} \rightarrow 4 \text{ Fe} + 3 \text{ CO}_2$   
What mass of iron will be produced when 50.0 g of iron(III) oxide react with an excess of carbon? (35.0 g)
- For the reaction:  $2 \text{ KClO}_{3(s)} \rightarrow 2 \text{ KCl}_{(s)} + 3 \text{ O}_{2(g)}$ 
  - What volume of oxygen gas, measured at STP, is formed by the reaction of 75.0 g of potassium chlorate? (20.6 L)
  - How many moles of potassium chlorate must react in order to form 2.5 g of potassium chloride? (0.034 mol)
- For the reaction:  $6 \text{ CO}_{2(g)} + 6 \text{ H}_2\text{O}_{(g)} \rightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6 \text{ O}_{2(g)}$ 
  - Which substance is in excess when 100 g of carbon dioxide and 50.0 g of water are reacted? (water)
  - How much glucose will be produced by the reaction in part (a)? (68.2 g)
- Octane burns according to the reaction:  $2 \text{ C}_8\text{H}_{18(l)} + 25 \text{ O}_{2(g)} \rightarrow 16 \text{ CO}_{2(g)} + 18 \text{ H}_2\text{O}_{(g)}$ 
  - What mass of carbon dioxide will be produced when 180 g of octane are completely burned? (555 g)
  - What volume of carbon dioxide, at STP, will be produced when 62.7 g of oxygen are completely reacted? (28.1 L)

# Review Answers Chemical Reactiona

**Camilla Rothe**



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