

# UNIT 1 NOTES

## Biological Chemistry Review

### Why study Chemistry in Biology?

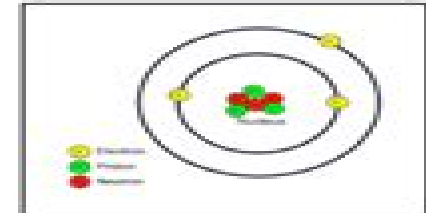
- The understanding of the basic chemistry of life is vital to a full comprehension of biology.
- Important atoms include:
  - Oxygen (O)
  - Carbon (C)
  - Hydrogen (H)
  - Nitrogen (N)
  - Phosphorous (P)
  - Potassium (K)
  - Sulphur (S)

### Atoms:

Atom = smallest part of an element that displays the properties of that element

- Made up of 3 subatomic particles:
  - Protons (positively charged)
  - Neutrons (uncharged / neutral)
  - Electrons (negatively charged) – found outside of the nucleus

Found in the nucleus



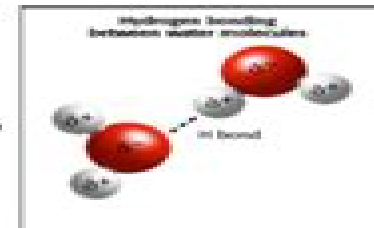
### Ions:

Ions = atoms that have gained or lost electrons

- Signified by either (+) or (-)
- Ex.  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Cl}^-$ ,  $\text{Ca}^{2+}$

### Types of bonding:

1. Ionic bonds:
  - Transfer of electrons from a metal (positive ion) to a non-metal (negative ion)
2. Covalent bonds:
  - Atoms share electrons between 2 non-metals
  - Can be single covalent bond (sharing 2 electrons), double covalent bond (sharing 4 electrons) or triple covalent bond (sharing 6 electrons)
    - Nonpolar covalent bonds
      - Electrons shared fairly equally between atoms
    - Polar covalent bonds
      - Unequal sharing of electrons
3. Hydrogen bonds:
  - When a covalently bonded H within a molecule (eg  $\text{H}_2\text{O}$ ) is also attracted to another atom (usually O, N or F) of a *different* molecule.
  - Always represented by dotted lines



### Inorganic Compounds:

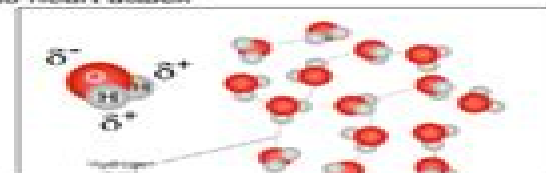
Inorganic Compounds	Organic Compounds
Usually metal + non-metal	<b>Always</b> contain carbon and hydrogen
Usually <i>ionic</i> bonding	<i>Always</i> covalent bonding
Associated with <b>non-living</b>	Associated with <b>living</b> things
Ex. Salts, water, acids, bases, buffers	Ex. Carbohydrates, lipids, proteins and nucleic acids

#### A. Salts:

- Needed for survival
- Too little → fatigue, exhaustion, cramps, nausea, thirst, decrease in blood pressure and even death
- Too much → dehydration of cells and hypertension; can lead to heart attack

#### B. Water:

- Structure of water:
- Biological importance is due to its chemical structure
- Water is a **polar** molecule
  - **How does polarity of water produce H-bonding?**
    - Oxygen is bigger & negatively charged
    - H is smaller & positively charged



# Review Biological Chemistry Section 2 Answer Sheet

**LP Steffe**



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