Chemistry Semester Exam Review

Complete all problems on notebook paper. Transfer your final answers to the answer sheet.

Measurements and Calculations

Significant Figures – adding, subtracting, multiplying, dividing Scientific Notation - adding, subtracting, multiplying, dividing Metric Conversions Dimensional Analysis Temperature Conversions Density – D=m/V, V=LWH, $V=0r^2h$, $V=V_r - V_s$

Determine the number of significant figures in the following numbers. c. 4 SE d. 6 SF

- a. 1 SF b. 2 SE 1. 0.0000787690 a
 - 2. 900000 km
 - 3. 0.009001 mm
 - 4. 9.100 cg

Record your answer with the correct number of significant figures and units.

- 5. 45.6 + 3.59a. 49.19 g
- b. 49.0 g
- c. 49.2 g
- d. 49.20 g

- 0.00134 mL 0.00023 mL
 - a. 0.00111 mL b. 11 mL
- c. 0.0011 mL d. 0.00157 mL

- 7. (56.2 cm)(1.234 cm)
 - a. 69.3 cm²
- b. 69.4 cm²
- c. 69.3508 cm² d. 69.35 cm²

e. 10 SE

- (3.20 km²)/(4.0 km) a. 13 km

- b. 8 km
- c. 0.80 km
- d. 2.56 km

Convert the following numbers from scientific notation to ordinary notation.

- 9. 3.02 x 10⁻³ q
 - a. .302 x 10⁻³ q
- b. 3.02 q c. 0.00302 q
- d. 3020 q

- 10. 5.791 x 10⁵ m
 - a. 579100 m
- b. 5.791 m c. 0.00005791 m d. 57910 m

Convert the following numbers from ordinary notation to scientific notation.

- 11.4560 cm
 - a. 456 x 10³ cm b. 4.56 x 10³ cm c. 4.56 x 10⁻³ cm d. 4.56 cm
- 12.0.0076 q
 - a. 7.6 x 10⁻³ g b. 7.6 x 10³ g c. 76 x 10⁻³ g d. 0.0076 x 10⁻³ g

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Steve Langfield, Dan Evans, Alison Wright

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