

REVIEW FOR THERMOCHEMISTRY TEST

1. The combustion of methanol is shown by the following equation:



- a) Using the data below, find the heat of reaction for the combustion of methanol.



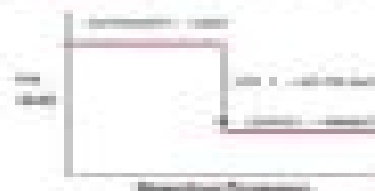
- b) What is the molar enthalpy of combustion for methanol?

$$\Delta H = -1278 \text{ kJ/mol} \text{ and } -639 \text{ kJ/mol}$$

- c) Express the energy change for the target reaction in the form $\Delta H_{\text{reaction}}$.

$$\Delta H_{\text{reaction CH}_3\text{OH}} = -639 \text{ kJ/mol}$$

- d) Draw the potential energy diagram to illustrate the energy change for the target reaction.



Review For Unit 12 Thermochemistry Test

Kaplan Test Prep



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