CHAPTER 9 REVIEW

Stoichiometry

MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation?

40,0 9/wolb. What is the molar mass of C3H4?

E. What is the mole ratio of O2 to H2O in the above equation?

O. 20 wol d. How many moles are in an 8.0 g sample of C₃H₄?

e. If z mol of C₃H₄ react, how many moles of CO₂ are produced, in terms of z? mol CO₂ = (2 mol) (3 co₂) = 3

2. a. What is meant by ideal conditions relative to stoichiometric calculations?

STP = 101.3 kPa

O°C

b. What function do ideal stoichiometric calculations serve?

Allow us to use known/constant values

e. Are actual yields typically larger or smaller than theoretical yields?

- somalley

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

3. Assume the reaction represented by the following equation goes all the way to completion:

 $N_2 + 3H_2 \rightarrow 2NH_2$

4 MOI H2 a. If 6 mol of H2 are consumed, how many moles of NH3 are produced? mol NH3 = (b mol Hz) 2 mol NH3 = 4 mol

8.5 g NH3 b. How many grams are in a sample of NH3 that contains 3.0 × 10²³ molecules?

gNH3= (3,0×1023 lemus)

602x1023mdrc

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