

CHAPTER 17

RADIOACTIVITY AND NUCLEAR CHEMISTRY

Practice Problems

1. (a) Write a nuclear equation for alpha decay of Po-210 .



- (b) Write a nuclear equation for the beta decay of Th-231 .



- (c) Write a nuclear equation for positron emission by Co-57 .



2. A 10.0-g sample of a radioactive nuclide is administered to a patient for cancer diagnosis. If the nuclide has a half-life of 1.5 hours, how much of the nuclide remains in the patient after 4.5 days?

$$\text{4.5 days} \times \frac{24 \text{ hours}}{1 \text{ day}} \times \frac{60 \text{ min}}{1 \text{ hour}} = 6480 \text{ min} \quad \frac{6480 \text{ min}}{90 \text{ min}} = 72 \text{ half-lives}$$

- (a) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (b) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (c) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (d) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (e) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (f) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (g) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (h) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (i) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (j) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$

3. A mammoth skeleton has a carbon-14 content of 18.3% of that found in living organisms. When did the mammoth live?

The age of the mammoth is 14,400 years. The half-life of C-14 is 5,730 years.

$$18.3\% = \left(\frac{1}{2}\right)^n \quad n = 2.5 \text{ half-lives}$$

$$2.5 \times 5,730 \text{ years} = 14,325 \text{ years}$$

Review Nuclear Chemistry Answer Key

**Andrew Burrows, John Holman, Simon
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 - The Rise of Digital Reading Review Nuclear Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Nuclear Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Nuclear Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Nuclear Chemistry Answer Key

- Personalized Recommendations
- Review Nuclear Chemistry Answer Key User Reviews and Ratings
- Review Nuclear Chemistry Answer Key and Bestseller Lists
- 5. Accessing Review Nuclear Chemistry Answer Key Free and Paid eBooks
 - Review Nuclear Chemistry Answer Key Public Domain eBooks
 - Review Nuclear Chemistry Answer Key eBook Subscription Services
 - Review Nuclear Chemistry Answer Key Budget-Friendly Options
- 6. Navigating Review Nuclear Chemistry Answer Key eBook Formats
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 - Review Nuclear Chemistry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Nuclear Chemistry Answer Key
 - Highlighting and Note-Taking Review Nuclear Chemistry Answer Key
 - Interactive Elements Review Nuclear Chemistry Answer Key
- 8. Staying Engaged with Review Nuclear Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Nuclear Chemistry Answer Key
- 9. Balancing eBooks and Physical Books Review Nuclear Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Nuclear Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Nuclear Chemistry Answer Key
 - Setting Reading Goals Review Nuclear Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Nuclear Chemistry Answer Key

- Fact-Checking eBook Content of Review Nuclear Chemistry Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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