

# Gas Laws Problem Packet

**Gas Laws Packet**

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

**Introduction**

22.4 L = 1 mol at STP

Unit	Standard Pressure
liters	22.4
atmos	1.0
temp	0
mol	1.0

1. All standard temperature and pressure (STP) are 0°C and 1.0 atm.

2. The gas laws:

Gas Law	Equation
Boyle's Law	$P_1V_1 = P_2V_2$
Charles's Law	$V_1/T_1 = V_2/T_2$
Gay-Lussac's Law	$P_1/T_1 = P_2/T_2$
Avogadro's Law	$V/n = k$
Combined Gas Law	$P_1V_1/T_1 = P_2V_2/T_2$
Standard Gas Law	$PV = nRT$

3. The constant  $R$  used in the standard gas law depends on the unit of pressure:

- liters  $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$
- atmos  $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

**Helpful Hints**

- Temperature must always be in Kelvin ( $0^\circ\text{C} + 273 = 273\text{K}$ )
- Units are given from different units of pressure, (liters) and
- Circle two columns before solving, initial and final. Write what you're given in the problem. How the law that uses the variables you're given, and solve for the unknown. Do that for each law, write what you are given, what those in equation, and solve for unknown.
- Avogadro's Law: When 1 mol of gas is at STP, it occupies 22.4 liters.
- Standard Gas Law: You will use this to solve the molecular mass for these problems. Divide and take what you have the rate of effusion and molecular masses in the context units before solving the equation (this is an inverted relationship). You will evaluate the inverse and it means both sides of the equation before moving things.

Charles's Law  $V_1/T_1 = V_2/T_2$

Boyle's Law  $P_1V_1 = P_2V_2$

The image shows a stack of papers with handwritten notes and equations. A yellow pencil is positioned diagonally across the papers. A grey calculator is placed in the foreground, partially overlapping the papers. The papers contain various gas law equations and problem-solving steps.

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# Review Packet Gas Laws Answer

**Vincent D. Kracum**



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