

DESCRIBING CHEMICAL REACTIONS

Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation

Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical ____1____. The substances that undergo a chemical change are the ____2____. The new substances formed in a chemical reaction are the ____3____. In accordance with the law of conservation of ____4____, a chemical equation must be balanced. When balancing an equation, you place ____5____ in front of reactants and products so that the same number of atoms of each ____6____ are on each side of the equation. An equation must never be balanced by changing the ____7____ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is ____8____; for a solid, ____9____; and for a gas, ____10____. A substance dissolved in water is designated ____11____. If a ____12____ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- _____ 14. The symbol Δ placed over the arrow in an equation means that heat is supplied to the reaction.
- _____ 15. Atoms are destroyed in a chemical reaction.
- _____ 16. A skeleton equation is not a balanced equation.

Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I
- Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.
 - Acetylene gas (C_2H_2) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

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Sandy Fritz



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1. Understanding the eBook Review And Reinforce Describing Chemical Reactions Answers
 - The Rise of Digital Reading Review And Reinforce Describing Chemical Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review And Reinforce Describing Chemical Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review And Reinforce Describing Chemical Reactions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review And Reinforce Describing Chemical Reactions Answers
 - Personalized Recommendations
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 - Review And Reinforce Describing Chemical Reactions Answers and Bestseller Lists
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 - Review And Reinforce Describing Chemical Reactions Answers Compatibility with Devices
 - Review And Reinforce Describing Chemical Reactions Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review And Reinforce Describing Chemical Reactions Answers
 - Highlighting and Note-Taking Review And Reinforce Describing Chemical Reactions Answers
 - Interactive Elements Review And Reinforce Describing Chemical Reactions Answers
 8. Staying Engaged with Review And Reinforce Describing Chemical Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review And Reinforce Describing Chemical Reactions Answers
 9. Balancing eBooks and Physical Books Review And Reinforce Describing Chemical Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review And Reinforce Describing Chemical Reactions Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Review And Reinforce Describing Chemical Reactions Answers
 - Setting Reading Goals Review And Reinforce Describing Chemical Reactions Answers
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Review And Reinforce Describing Chemical Reactions Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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