

# 14-1 Review and Reinforcement

## Condensed States of Matter

Fill in the following chart comparing solids and liquids.

Physical Property	Liquids	Solids
1. Compressibility		
2. Density		
3. Shape		
4. Rate of Diffusion		

Answer each of the following questions in the space provided.

5. Describe the behavior of the molecules in a liquid. Explain this behavior in terms of intermolecular forces.
- \_\_\_\_\_
- \_\_\_\_\_
6. How do intermolecular forces determine molecular arrangement in solids?
- \_\_\_\_\_
- \_\_\_\_\_

If the statement is true, write "true." If it is false, change the underlined word or words to make the statement true. Write your answer on the line provided.

- \_\_\_\_\_ 7. Intermolecular forces determine metallic properties such as the boiling point of a substance.
- \_\_\_\_\_ 8. Intermolecular forces are forces of attraction between atoms.
- \_\_\_\_\_ 9. Chemical bonds are the forces between molecules.
- \_\_\_\_\_ 10. Covalent bonds result from electrons being shared between atoms in a molecule.
- \_\_\_\_\_ 11. Intermolecular forces result from the electron interactions between neighboring molecules.
- \_\_\_\_\_ 12. When the electron cloud is not distributed symmetrically, a molecule has a polarity.
- \_\_\_\_\_ 13. Chemical bonds include ionic bonds, metallic bonds, and atomic bonds.
- \_\_\_\_\_ 14. Intermolecular forces include dispersion forces, dipole-dipole forces, and helium-bond forces.

# Review And Reinforcement Answers 14 1

**Guillaume Favre**



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