

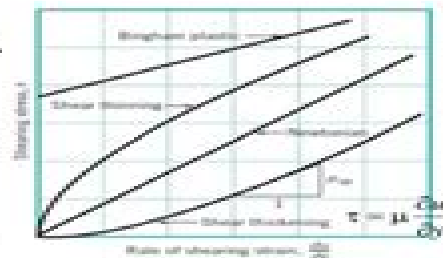
Ideal Gases

- From chemistry: $PV = nRT$ (V is volume)
 - $n = m / M$ is the number of moles
 - for mass in kg n is in kilogram moles (kmol); for mass in lb_{mass}, n is in pound moles (lbmol)
 - $R = 8.31447 \text{ kJ/kmol}\cdot\text{K} = 10.7316 \text{ psia}\cdot\text{ft}^3 / \text{lbmol}\cdot\text{R}$ is universal gas constant
 - $R = R/M$ is engineering gas constant that is different for each gas
 - Real gases like ideal gases at low pressures
 - $P = nRT / V = (m/M)RT / V = (m/V)(R/M)T$
- $P = \rho RT$

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Viscosity

Newtonian and non-Newtonian
Variation of shear stress with rate of shearing strain for several types of fluids, including common non-Newtonian fluids.



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Surface Tension Rise



- Vertical force balance: $\gamma \pi R^2 h = 2\pi R \sigma \cos \theta$
 - Surface tension depends on fluid and temperature, wetting angle, θ , depends on fluid and surface
- $$h = \frac{2\sigma \cos \theta}{\gamma R}$$

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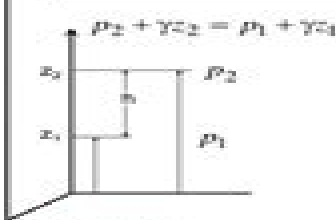
Pressure Relations

- Pressure is a scalar
- The force exerted by a pressure is the same in all directions
- Want to see how pressure changes in a static (nonmoving fluid)
- Look at balance of pressure force and fluid weight over a differential volume element, $\Delta x \Delta y \Delta z$

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Incompressible Fluid



- Which pressure is higher?
 $p_1 = p_2 + \gamma(z_2 + z_1)$
 $p_1 = p_2 + \gamma h > p_2$
- Pressure increases with depth

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Gage and Absolute Pressure



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Review For Final Exam Fluid Mechanics

G Thomas



Review For Final Exam Fluid Mechanics:

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