



## 2

## ATOMIC THEORY: ISOTOPES

Solutions

Complete the following table.

Isotope Name		Atomic Number	Mass Number	Symbol	Number of Protons	Number of Neutrons
1.	carbon-14	6	14	$^{14}_6\text{C}$	6	8
2.	oxygen-16	8	16	$^{16}_8\text{O}$	8	8
3.	polonium-112	84	212	$^{212}_{84}\text{Po}$	84	128
4.	uranium-238	92	238	$^{238}_{92}\text{U}$	92	146
5.	hydrogen-2	1	2	$^2_1\text{H}$	1	1
6.	helium-4	2	4	$^4_2\text{He}$	2	2
7.	thorium-232	90	232	$^{232}_{90}\text{Th}$	90	142
8.	carbon-12	6	12	$^{12}_6\text{C}$	6	6
9.	lawrencium-257	103	257	$^{257}_{103}\text{Lr}$	103	154
10.	hydrogen-1	1	1	$^1_1\text{H}$	1	0

# Review Sheet Calculations Test 2 Answer Key

**RJ Alexander**



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