

Review Sheet – Chemistry, Level 3 – Ch. 14: Liquids & Solids

Name: _____ ANSWERS _____

Date: _____

Period: _____

This test covers materials in the textbook in Chapters 14 (pp. 456–491). You are always responsible for the topics on the previous tests.

Test Expectations

The student should be able:

- *Describe how the Kinetic-Molecular Theory of Matter predicts the properties of substances*
- *Analyze the forces and relative strengths of the forces between particles in a solid, liquid or gas and compare the substances in terms of basic properties such as melting point, boiling point, viscosity, surface tension, or vapor pressure*
- *Describe the type and properties of one of the four solids*
- *Draw or use phase diagrams or heating curves to describe phase changes*

1. Define each of the following terms clearly and completely.

(a) condensed state of matter –

(b) phase –

(c) phase diagram –

(d) heating curve –

(e) intramolecular force –

(f) intermolecular force –

(g) metallic bond –

(h) hydrogen bond –

(i) dipole-dipole force –

(j) dispersion force –

(k) permanent dipole –

(l) temporary dipole –

(m) induced dipole –

(n) viscosity –

(o) surface tension –

(p) water –

(r) crystalline solid –

(s) amorphous solid

(t) adhesion –

(u) cohesion –

(v) covalent-network solid –

(w) phase change –

(x) vapor pressure –

(y) heat of fusion –

(z) heat of vaporization –

(aa) sublimation –

(bb) deposition –

(cc) vaporization –

(dd) condensation –

(ee) phase change –

(ff) freezing –

(gg) melting –

(hh) Kinetic-Molecular Theory of Matter –

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