

## Compounds

| Word                             | Definition  |
|----------------------------------|---|
| Binary compound                  | A compound that consists of two elements.   |
| Dipole attractions               | The attraction of the $\delta^-$ end of one polar molecule to the $\delta^+$ end of another polar molecule.   |
| Dipole moment                    | An arrow along the line of symmetry in a polar molecule that shows the net direction that electrons are being pulled towards the partially negative end of the molecule.            |
| Electronegativity difference     | The difference in electronegativity between two elements in a bond.   |
| Electrical conductivity (metals) | The ability of a substance to allow electrons to pass from atom to atom through itself from a source or electricity to an electrical ground.  |
| Electrolyte                      | A solution containing dissolved ions that can conduct electricity through it.   |
| Empirical formula                | The simplest whole-number mole ratio of elements in a compound, used to write the formulas of ionic compounds.  |
| Formula mass                     | The sum of the atomic masses of an element or compound, measured in grams/mole. Reported to the nearest tenth of a gram per mole.   |
| Hydrogen bonds                   | The strong attraction of the H ( $\delta^+$ ) end of one polar molecule to the N, O or F ( $\delta^-$ ) end of another polar molecule. The two ends form temporary covalent bonds.  |
| Intermolecular Attractive Force  | The forces that hold molecules together in the solid and liquid phase. These are the forces that must be overcome to melt or boil a substance. Also called "van der Waal's forces". |
| Ionic compound                   | Compounds consisting of a metal and a nonmetal that are ionically bonded in a whole-number ratio.   |
| London dispersion force          | The weak attractive force caused by temporary dipoles in nonpolar molecules.  |
| Metallic bond                    | A bond formed between metal atoms as they collectively share their conducting electrons evenly between metal kernels.   |
| Molecular formula                | The actual number of nonmetal atoms in a molecule, a whole-number multiple of the empirical formula.  |
| Molecule                         | A particle made of nonmetal atoms covalently bonded together to form a distinct particle.   |
| Network solid                    | A crystal lattice formed from covalently bonded nonmetal atoms with no distinct molecules.  |
| Nonpolar molecule                | A molecule with symmetrical electron distribution resulting in any polar bonds canceling each other out to yield no partially charged ends.   |
| Percent composition              | The formula mass of an element divided by the formula mass of the compound that contains the element multiplied by 100.   |
| Polar molecule                   | A molecule with asymmetrical electron distribution resulting in partially charged ends.   |
| Polyatomic ion                   | An ion formed by atoms bonding together in such a way that a net charge is formed.  |
| Ternary compound                 | A compound that consists of three or more elements, usually containing a polyatomic ion.  |

# Rosengaten Unit 7 Answers

**N Noddings**



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