



Restoring Balance

LeChâtelier's Principle and Equilibrium

Introduction

Cheroical equilibrium is a true balancing as: What happens when the balance is disturbed? The purpose of this lab is to observe the effects of concentration and temperature on equilibrium and to visualize how balance can be restored based on LeChitelier's Principle.

Concepts

· Chemical equilibrium

· LeChäselier's principle

Complex-ton residues

Exactlements via enclustrements reactions.

Background

Not, all chemical reactions proceed to completion, giving 100% yield of processes. In fact, most, chemical reactions are reversible, meaning they can go both ways. In the forecast direction, reactions interact to make process while in the reverse direction, the products revers both to reactions. This idea is represented symbolically using disable arrows as shown for the reversible reaction of incline nucleonless and indice into the give principle into the term of the control of the process of the

$$T_i(sc) + T(sc) e^{-s}T_i(sc)$$
 Equation 2

In a closest systems, any revenible reaction will eventually reach a system is sixturned between the forward and revenue reactions. A system is said as reach chemical equilibrium when the rate of the forward reaction equals the rate of the revenue reaction. As this point, no further changes will be observed in the ansounce of either the reactions or products. Chemical equilibrium can be further defined, therefore, as the state at which the concentrations of reactions and products remain constant with time. This does not mean the concentrations of reactions and products are equal. The forward and revenue reactions create an equal halance of opposing rates.

What happens when the balance is disturbed when to the addition of more reactants or products or due to changes in the temperature or pressure? LeChitelier's Principle predicts how equilibrium can be testered:

> "Af an equilibrium system is subjected to a stress, He system will reset in such a way as to reduce the stress."

Any change that is made to a system at equilibrium is considered a stream—this includes adding or removing respents or changing the tempersure. To reduce the stress, one of two things can happen. A reversible reaction can disf; in the forward direction and make room products, thus using up some of the rescures. Alternatively, the reaction can shift in the reverse direction and re-form the reactarits, thus using up some of the products. The effect of temperature on a system at exhibitions depends on whether a reaction is endothermic (absorbs heat) or exothermic (produces heat). If a reaction is endothermic, best appears on the reactant side in the chemical equation. Increasing the temperature of an endorthermic reaction shifts the equifibrium in the forested direction to consume some of the excess energy and make more products. The opposite effect is observed for exothermic reactions. In the case of an exothermic reaction, heat appears on the product side in the chemical equation.

There asing the temperature of an exothermic reaction while the equilibrium in the reverse direction.

Experiment Overview

The purpose of this experiment is to investigate the effect of reaction conditions on the revenible formation of collab complex ions. When collab chloride breakydrate (CoCl₂%H₂O) is disolved in ethyl alcohol, there different solute species are present: Co^2 cations Cl anions and water molecules. There can react to form two different complex ions, $Co(H_2O)_{ij}^{-1}$, where the collab tion is surmanded by four chloride total and $Co(H_2O)_{ij}^{-1}$, in which the metal ton is surmanded by four chloride total.

Pre-Lab Questions

Indine (I_j) is only sparingly soluble in water (Equation 2). In the
presence of potassium todicie, a source of todicie (I_j) ions, todine
reacts to form trincitie (I_j) ions (Equation 3).

$$I_{p}(ac) + I(ac) e^{ac}I_{p}(ac)$$
 Equation 3

Use LeChbrelier's Principle to explain why the solubility of indire inwater increases as the concentration of patassium indice increases.

 Disselving aromonium abloride in water is an endothermic reaction (Research 4). The Le Châtelier's Principle to predict whether aromonium chloride will be more soluble in boton cold water.

Restoring Balance Equilibrium Lab Flinn Answers

W.J. Fellner

Restoring Balance Equilibrium Lab Flinn Answers: Maintaining and Restoring Balance W.J. Fellner,

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