

# Gravimetric Analysis of Chloride: Lab 4

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Section: 2A38

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## Introduction:

Gravimetric analysis is a technique designed to isolate and determine the measurement of mass of the ion. By separating and determining that mass of the ion in the solution we can determine the molar mass and the other molecule used in the compound ( $M_xCl_y$ ). Through gravimetric analysis the values of  $M$ ,  $x$ , and  $y$  can be determined. Typical gravimetric analysis has chemists weight the sample of the unknown substance, dissolve the sample in a solvent, add a reagent to form a precipitate, filter and dry the precipitate and calculate the percent of substance from the weight of the precipitate and the sample. For example, in this experiment the unknown compound of  $Cl$  is added to water, to ionize. Then mixed with both  $HNO_3$  and  $AgNO_3$ , which forms a precipitate through the reaction in **Figure 1.A**. After the precipitate is formed it is filtered and dried in a drying oven, after the final mass of  $AgCl$  is determined and used in the calculations to determine the mass of the compound  $M$ .



## Experimental:

Procedure: (found in Green Lab Manual Zhao, *General Chemistry Lab Manual* (1st Edition), pgs. 36-38)

Original Procedure: No Altered Procedure

## Data Analysis:

Gravimetric Analysis Data Collection

	Trial 1	Trial 2
Mass of $Cl$ sample:	0.2511 g	0.2500 g
Mass of crucible:	32.2227 g	32.1936 g
Mass of crucible + precipitate:	32.8383 g	32.7750 g
Mass of $AgCl$ precipitate:	0.6156 g	0.5814 g

Trial 1:

Trial 2:

# Results For Gravimetric Analysis Lab Report

**James Sherwood Fritz, George H.  
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