- Given the nuclear equation: <sup>14</sup>7N + X → <sup>16</sup>3 O + <sup>2</sup>1H
   What is particle X?
- (A) an alpha particle
  (C) a deuteron
  (D) a triton
- The nucleus of a radium-226 atom is unstable, which causes the nucleus to spontaneously
  - A) absorb electrons B) absorb protons C) decay D) oxidize
- Which group of nuclear emissions is listed in order of increasing charge?
  - A) alpha particle, beta particle, gamma radiation -
  - B) gamma radiation, alpha particle, beta particle -
  - positron, alpha particle, neutron
  - D) neutron, positron, alpha particle
- 4. Which list of nuclear emissions is arranged in order from the *least* penetrating power to the greatest penetrating power?
  - A) alpha particle, beta particle, gamma ray
  - B) alpha particle, gamma ray, beta particle
  - C) gamma ray, beta particle, alpha particle
  - D) beta particle, alpha particle, gamma ray
- Base your answer to the following question on Given the nuclear equation:

$$^{1}iH + X \rightarrow ^{6}iLi + ^{4}He$$

The particle represented by X is

- A) %4Li B) %4Be C) 105Be D) 106C
- 6. Positrons are spontaneously emitted from the nuclei of
  - A) potassium-37
- B) radium-226
- C) nitrogen-16 D) thorium-232
- What is the decay mode of <sup>37</sup>K?
- A) B B) B+ C) 7 D) a
- 8. Which equation represents positron decay?
  - A)  $^{27}$  37 Rb  $\rightarrow$   $^{0}$  -1 e +  $^{27}$  38 Sr
  - B)  $^{277}$  92 U  $\rightarrow$   $^{223}$  90 Th  $+ ^{4}$  2 He
  - ©  $^{27}_{13}Al + ^{4}_{2}He \rightarrow ^{30}_{15}P + ^{1}_{0}n$
  - D) 116C -0+1e+115B

- As <sup>14</sup>C decays to <sup>14</sup>N, the number of protons in the nucleus
  - A) decreases

- B) increases
- C) remains the same
- 10. Atoms of I-131 spontaneously decay when the
  - A) stable nuclei emit alpha particles
  - B) stable nuclei emit beta particles
  - C) unstable nuclei emit alpha particles
  - D) unstable nuclei emit beta particles
- 11. Which statement best describes gamma radiation?
  - A) It has a mass of 1 and a charge of 1.
  - B) It has a mass of 0 and a charge of -1.
  - C) It has a mass of 0 and a charge of 0.
  - D) It has a mass of 4 and a charge of +2.
- 12. Which nuclear decay emission consists of energy, only?
  - A) alpha particle
- B) beta particle
- C) gamma radiation
- D) positron
- 13. Which kind of nuclear radiation has high energy and no mass?
  - A) alpha

B) beta

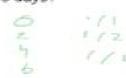
C) gamma

- D) neutron
- 14. The half-life of a radioactive isotope is 20.0 minutes. What is the total amount of a 1.00-gram sample of this isotope remaining after 1.00 hour?
  - A) 0.500 g

B) 0.333 g

C) 0.250 g

- D) 0.125 g
- 15 A radioactive element has a half-life of 2 days. Which fraction represents the amount of an original sample of this element remaining after 6 days?
  - 70 J
  - B)  $\frac{1}{2}$
  - $\bigcirc$   $\frac{1}{3}$
  - D) <u>1</u>



# **Review Nuclear Chemistry Section Answers**

**D** Kirk

**Review Nuclear Chemistry Section Answers:** 

Whispering the Secrets of Language: An Emotional Journey through Review Nuclear Chemistry Section Answers

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