

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
Conducts electricity Contains hydrogen ion Proton donor Turns litmus paper red React with most metals Releases hydronium Ex: lemon Vinegar Battery acid (sulfuric acid)	conducts electricity contains hydroxide ion proton acceptor turns litmus paper blue releases hydroxide ion Ex: Windex bleach soap

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, **acids** or **bases**:

- A** produce hydrogen ions (H^+) in solution
- B** have a bitter taste
- B** produce hydroxide ions (OH^-) ions
- A** react with bases to form salts
- A** have a sour taste
- B** have a slippery, soapy feel
- B** react with acids to form salts
- A** react with many metals

3. Write an equation for the dissociation of the following **acids** in water: follow example

- $HClO_4 \rightarrow H^+ + ClO_4^-$
- $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$
- $HC_2H_3O_2 \rightarrow H^+ + C_2H_3O_2^-$
- $H_2S \rightarrow 2H^+ + S^{2-}$
- $HCl \rightarrow H^+ + Cl^-$
- $HNO_3 \rightarrow H^+ + NO_3^-$

4. Write an equation for the dissociation (ionization) of the following **bases** in water: (follow example)

- $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$
- $NaOH \rightarrow Na^+ + OH^-$
- $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$
- $KOH \rightarrow K^+ + OH^-$
- $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$
- $LiOH \rightarrow Li^+ + OH^-$

5) A. What type of reaction is a **neutralization** reaction? **Double replacement**

B. Write the net equation for neutralization:



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Table of Contents Review And Reinforcement Acids And Bases Answers

1. Understanding the eBook Review And Reinforcement Acids And Bases Answers
 - The Rise of Digital Reading Review And Reinforcement Acids And Bases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review And Reinforcement Acids And Bases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review And Reinforcement Acids And Bases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review And Reinforcement Acids And Bases Answers
 - Personalized Recommendations
 - Review And Reinforcement Acids And Bases Answers User Reviews and Ratings
 - Review And Reinforcement Acids And Bases Answers and Bestseller Lists
5. Accessing Review And Reinforcement Acids And Bases Answers Free and Paid eBooks

- Review And Reinforcement Acids And Bases Answers Public Domain eBooks
 - Review And Reinforcement Acids And Bases Answers eBook Subscription Services
 - Review And Reinforcement Acids And Bases Answers Budget-Friendly Options
6. Navigating Review And Reinforcement Acids And Bases Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Review And Reinforcement Acids And Bases Answers Compatibility with Devices
 - Review And Reinforcement Acids And Bases Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review And Reinforcement Acids And Bases Answers
 - Highlighting and Note-Taking Review And Reinforcement Acids And Bases Answers
 - Interactive Elements Review And Reinforcement Acids And Bases Answers
 8. Staying Engaged with Review And Reinforcement Acids And Bases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review And Reinforcement Acids And Bases Answers
 9. Balancing eBooks and Physical Books Review And Reinforcement Acids And Bases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review And Reinforcement Acids And Bases Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Review And Reinforcement Acids And Bases Answers
 - Setting Reading Goals Review And Reinforcement Acids And Bases Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Review And Reinforcement Acids And Bases Answers
 - Fact-Checking eBook Content of Review And Reinforcement Acids And Bases Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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from the raw material supplier, which ... to conduct a mass balance test every 6 months for each claim or a single mass balance test every.