CHAPTER 6 REVIEW

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- _A_ The notation for sodium chloride, NaCl, stands for one
 (a)formula unit.
 (b)molecule.(d)

 crystal.
 atom.
- _D_ In a crystal of an ionic compound, each cation is surrounded by a number of
 (a)molecules.
 (c) dipoles.
- (a)molecules. (c) dipoles. (b)positive ions. (d) negative ions.
- B Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
- (a)higher in potential energy. (c) equal in potential energy.
- (b)lower in potential energy. (d) unstable.
- B The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a)Compound A is not an ionic compound.
- (b)It will be more difficult to break the bonds in compound A than those in compound B.
- (c)Compound B has larger crystals than compound A.
- (d)Compound A has larger crystals than compound B.
- B The forces of attraction between molecules in a molecular compound are generally
- (a)stronger than the attractive forces among formula units in ionic bonding.
- (b)weaker than the attractive forces among formula units in ionic bonding.
- (c)approximately equal to the attractive forces among formula units in ionic bonding.
- (d)equal to zero.
- 6.Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

7. What type of energy best represents the strength of an ionic bond?

Lattice energy

8.What types of bonds are present in an ionic compound that contains a polyatomic ion?

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Raffaela Di Napoli

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- 1. Understanding the eBook Review Chemical Bonding Section 5 Answer Key
 - The Rise of Digital Reading Review Chemical Bonding Section 5 Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Review Chemical Bonding Section 5 Answer Key
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Chemical Bonding Section 5 Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Review Chemical Bonding Section 5 Answer Key
 - Personalized Recommendations
 - Review Chemical Bonding Section 5 Answer Key User Reviews and Ratings
 - Review Chemical Bonding Section 5 Answer Key and Bestseller Lists
- 5. Accessing Review Chemical Bonding Section 5 Answer Key Free and Paid eBooks
 - Review Chemical Bonding Section 5 Answer Key Public Domain eBooks
 - Review Chemical Bonding Section 5 Answer Key eBook Subscription Services
 - Review Chemical Bonding Section 5 Answer Key Budget-Friendly Options
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 - Review Chemical Bonding Section 5 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Chemical Bonding Section 5 Answer Key
 - Highlighting and Note-Taking Review Chemical Bonding Section 5 Answer Key
 - Interactive Elements Review Chemical Bonding Section 5 Answer Key
- 8. Staying Engaged with Review Chemical Bonding Section 5 Answer Key

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Review Chemical Bonding Section 5 Answer Key
- 9. Balancing eBooks and Physical Books Review Chemical Bonding Section 5 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Chemical Bonding Section 5 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Chemical Bonding Section 5 Answer Key
 - Setting Reading Goals Review Chemical Bonding Section 5 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Chemical Bonding Section 5 Answer Key
 - Fact-Checking eBook Content of Review Chemical Bonding Section 5 Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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